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6. If there are too few pairs to give each atom eight electrons, change the single bonds between two atoms to either double or triple bonds by moving the unbonded pairs of electrons next to a bonding pair.

Once the octet rule has been satisfied for each atom in the molecule then you may replace each pair of dots between two atoms with a dash.

- 8. Now check your structure by
- a) count the total number of electrons to make sure you did not lose or gain electrons during the process.
- b) Use FORMAL CHARGE (FC) calculations as a guideline to the correct structure. A zero formal charge is usually a good indication of a stable structure.

FC (X) = # of valence electrons - (1/2 bonding electrons + nonbonding electrons)

For our example:
$$FC(C) = 4 - (1/2 + 0) = 0$$

 $FC(0) = 6 - (1/2 + 4) = 0$

Guide Line To Provide Good Bonding

Michael Brown

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