

- Q.20 For a reversible reaction forward and backward rate constant respectively are  $2.5 \times 10^{-4}$  and  $7.5 \times 10^{-4}$ . What is the value of equilibrium constant [RPMT-2001]  
 (1) 3 (2)  $1/3$  (3) 1 (4) None
- Q.21 If the equilibrium constant for the reaction  $N_2 + 3H_2 \rightarrow 2NH_3$  is K then the equilibrium constant for the reaction  $2N_2 + 6H_2 \rightarrow 4NH_3$  would be equal to [RPMT-2002]  
 (1)  $K^2$  (2)  $\sqrt{K}$  (3)  $\frac{1}{\sqrt{K}}$  (4)  $\frac{1}{K^2}$
- Q.22 The reaction quotient (Q) for the reaction  $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$  is given by  $Q = \frac{[NH_3]^2}{[N_2][H_2]^3}$ . The reaction will proceed from right to left if [CBSE-PMT-2003]  
 (1)  $Q < K_e$  (2)  $Q > K_e$  (3)  $Q = 0$  (4)  $Q = K_e$
- Q.23 The following equilibria are given [CBSE-PMT-2003]  
 $N_2 + 3H_2 \rightarrow 2NH_3$  .....  $K_1$   
 $N_2 + O_2 \rightarrow 2NO$  .....  $K_2$   
 $H_2 + \frac{1}{2}O_2 \rightarrow H_2O$  .....  $K_3$   
 The equilibrium constant of the reaction  $2NH_3 + \frac{5}{2}O_2 \rightarrow 2NO + 3H_2O$ , in terms of  $K_1$ ,  $K_2$  and  $K_3$  is :  
 (1)  $K_1 K_2 / K_3$  (2)  $K_1 K_3^3 / K_2$  (3)  $K_2 K_3^3 / K_1$  (4)  $K_1 K_2 K_3$
- Q.24 For the gas phase reaction,  
 $C_2H_4 + H_2 \rightleftharpoons C_2H_6$  ( $\Delta H = -32.7$  kcal),  
 carried out in a vessel, the equilibrium conc. of  $C_2H_4$  can be increased by [AMU 2003]  
 (1) Decreasing the pressure (2) Increasing the temperature  
 (3) Removing some  $C_2H_6$  (4) Adding some  $H_2$
- Q.25 Consider the following reversible reaction at equilibrium,  $2H_2O(g) \rightleftharpoons 2H_2(g) + O_2(g)$   
 $\Delta H = 24.7$  kJ. Which one of the following changes in conditions will lead to maximum decomposition of  $H_2O(g)$ ? [Kerala CEE 2004]  
 (1) Increasing both temperature and pressure  
 (2) Decreasing temperature and increasing pressure  
 (3) Increasing temperature and decreasing pressure  
 (4) Increasing temperature at constant pressure
- Q.26 In the reaction,  $A_2(g) + 4B_2(g) \rightleftharpoons 2AB_4(g)$   $\Delta H < 0$ , the formation of  $AB_4(g)$  will be favoured at [MP PMT 2004]  
 (1) High temperature, low pressure (2) Low temperature, high pressure  
 (3) High temperature, high pressure (4) Low temperature, low pressure

# Environntal Chemistry Mcq

**Julia Schneider**



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