

Empirical and Molecular Formulas Practice

Read the question carefully for the data provided. Answers must be shown for credit.

Find empirical formula:

1. A compound is found to contain 40.0 g of carbon and 6.67 g of hydrogen. Find the empirical formula.

$$40.0 \text{ g C} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 3.33 \text{ mol C}$$

$$6.67 \text{ g H} \times \frac{1 \text{ mol H}}{1.01 \text{ g H}} = 6.60 \text{ mol H}$$

C₃H₆

2. A compound is found to contain 20.0 g of carbon, 26.7 g of oxygen, and 6.67 g of hydrogen. Find the empirical formula.

$$20.0 \text{ g C} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 1.67 \text{ mol C}$$

$$26.7 \text{ g O} \times \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 1.67 \text{ mol O}$$



$$1.67 \text{ mol C} \times \frac{1 \text{ mol}}{1.67 \text{ mol}} = 1.00 \text{ mol} \quad 1.67 \text{ mol O} \times \frac{1 \text{ mol}}{1.67 \text{ mol}} = 1.00 \text{ mol} \quad 6.60 \text{ mol H} \times \frac{1 \text{ mol}}{1.67 \text{ mol}} = 4.00 \text{ mol H}$$

3. A 10.0 g sample of a compound, a gas, contains 2.00 g of carbon, 2.00 g of hydrogen, and 6.00 g of oxygen. Find the empirical formula.

$$2.00 \text{ g C} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 0.167 \text{ mol C}$$



$$6.00 \text{ g O} \times \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 0.375 \text{ mol O} \quad 0.167 \text{ mol C} \times \frac{1 \text{ mol}}{0.167 \text{ mol}} = 1.00 \text{ mol C}$$

$$1.98 \text{ mol H} \times \frac{1 \text{ mol}}{0.167 \text{ mol}} = 11.8 \text{ mol H} \quad 0.375 \text{ mol O} \times \frac{1 \text{ mol}}{0.167 \text{ mol}} = 2.25 \text{ mol O}$$

4. A 10.0 g sample of an unknown compound contains 2.00 g carbon, 2.00 g hydrogen, and 6.00 g oxygen. Find the empirical formula.

$$2.00 \text{ g C} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 0.167 \text{ mol C}$$



$$6.00 \text{ g O} \times \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 0.375 \text{ mol O}$$

$$0.167 \text{ mol C} \times \frac{1 \text{ mol}}{0.167 \text{ mol}} = 1.00 \text{ mol} \quad 1.98 \text{ mol H} \times \frac{1 \text{ mol}}{0.167 \text{ mol}} = 11.8 \text{ mol H} \quad 0.375 \text{ mol O} \times \frac{1 \text{ mol}}{0.167 \text{ mol}} = 2.25 \text{ mol O}$$

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