

# Empirical and Molecular Formulas Practice

Read the question carefully for the data provided. Answers must be shown for credit.

Find empirical formula:

1. A compound is found to contain 40.0 g of carbon and 6.67 g of hydrogen. Find the empirical formula.

$$40.0 \text{ g C} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 3.33 \text{ mol C}$$

$$6.67 \text{ g H} \times \frac{1 \text{ mol H}}{1.01 \text{ g H}} = 6.60 \text{ mol H}$$

C<sub>3</sub>H<sub>6</sub>

2. A compound is found to contain 20.0 g of carbon, 26.7 g of oxygen, and 6.67 g of hydrogen. Find the empirical formula.

$$20.0 \text{ g C} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 1.67 \text{ mol C}$$

$$26.7 \text{ g O} \times \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 1.67 \text{ mol O}$$



$$6.67 \text{ g H} \times \frac{1 \text{ mol H}}{1.01 \text{ g H}} = 6.60 \text{ mol H} \quad \frac{6.60 \text{ mol H}}{1.67 \text{ mol C}} = \text{C}_1\text{H}_4\text{O}_1$$

3. A 10.0 g sample of a compound, a gas, contains 2.50 g carbon, 2.50 g hydrogen, and 5.00 g oxygen. Find the empirical formula.

$$2.50 \text{ g C} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 0.208 \text{ mol C}$$



$$2.50 \text{ g H} \times \frac{1 \text{ mol H}}{1.01 \text{ g H}} = 2.48 \text{ mol H} \quad \frac{2.48 \text{ mol H}}{0.208 \text{ mol C}} = \text{C}_1\text{H}_{12}$$

$$5.00 \text{ g O} \times \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 0.312 \text{ mol O} \quad \frac{0.312 \text{ mol O}}{0.208 \text{ mol C}} = \text{C}_1\text{H}_{12}\text{O}_1$$

4. A 20.00 g sample of an unknown compound contains 12.00 g carbon, 2.00 g hydrogen, and 6.00 g oxygen. Find the empirical formula.

$$12.00 \text{ g C} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 1.00 \text{ mol C}$$



$$2.00 \text{ g H} \times \frac{1 \text{ mol H}}{1.01 \text{ g H}} = 1.98 \text{ mol H}$$

$$6.00 \text{ g O} \times \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 0.375 \text{ mol O} \quad \frac{0.375 \text{ mol O}}{0.375 \text{ mol C}} = \text{C}_1\text{H}_4\text{O}_1$$

# Empirical Formula Pogil Answers

**Scott C. Dulebohn**



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buffers typically consist of an acid base pair with the acid and base differing by the presence or absence of a proton a conjugate acid base pair for instance one of the buffers that maintain the ph of human blood involves carbonic acid  $\text{H}_2\text{CO}_3$  and its conjugate base the bicarbonate ion  $\text{HCO}_3^-$

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the reaction of an acid and a base is called a neutralization reaction although acids and bases have their own unique chemistries the acid and base cancel each other s chemistry to produce a rather innocuous substance water in fact the general reaction between an acid and a base is ce acid base water salt nonumber

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may 20 2018 acids react with bases to produce a salt compound and water when equal moles of an acid and a base are

combined the acid is neutralized by the base the products of this reaction are an ionic compound which is labeled as a salt and water

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nov 13 2022 suggest simple tests you could carry out to determine if an unknown substance is an acid or a base state the chemical definitions of an acid and a base in terms of their behavior in water write the formula of the salt formed when a given acid and base are combined