

Empirical and Molecular Formulas Practice

Read the question carefully for the data provided. Answers must be shown for credit.

Find empirical formula:

1. A compound is found to contain 40.94 g of iron and 35.48 g of sulfur. Find the empirical formula.

$$\text{40.94 g Fe} \times \frac{1 \text{ mol Fe}}{55.85 \text{ g Fe}} = 0.7332 \text{ mol Fe}$$

$$\text{35.48 g S} \times \frac{1 \text{ mol S}}{32.06 \text{ g S}} = 1.1066 \text{ mol S}$$

FeS

2. A compound is found to contain 20.0 g of carbon, 26.7 g of oxygen, and 6.7 g of hydrogen. Find the empirical formula.

$$\text{20.0 g C} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 1.66 \text{ mol C}$$

$$\text{26.7 g O} \times \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 1.67 \text{ mol O}$$



$$\text{6.7 g H} \times \frac{1 \text{ mol H}}{1.008 \text{ g H}} = 6.64 \text{ mol H} \quad \text{C}_1\text{O}_1\text{H}_6$$

3. A white solid is found to contain 40.94 g of iron, 35.48 g of sulfur, 20.0 g of carbon, 26.7 g of oxygen, and 6.7 g of hydrogen. Find the empirical formula.

$$\text{40.94 g Fe} \times \frac{1 \text{ mol Fe}}{55.85 \text{ g Fe}} = 0.7332 \text{ mol Fe}$$



$$\text{35.48 g S} \times \frac{1 \text{ mol S}}{32.06 \text{ g S}} = 1.1066 \text{ mol S} \quad \text{Fe-C}_1\text{O}_1\text{H}_6\text{S}_1$$

$$\text{20.0 g C} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 1.66 \text{ mol C} \quad \text{Fe-C}_1\text{O}_1\text{H}_6\text{S}_1\text{C}_2$$

4. A 20.00 g compound is analyzed and found to contain 12.00 g carbon, 13.33 g oxygen, and 4.67 g hydrogen. Find the empirical formula.

$$\text{12.00 g C} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 1.00 \text{ mol C}$$



$$\text{13.33 g O} \times \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 0.833 \text{ mol O}$$

$$\text{4.67 g H} \times \frac{1 \text{ mol H}}{1.008 \text{ g H}} = 4.64 \text{ mol H} \quad \text{H}_2\text{O-C}_1\text{O}_1\text{H}_2$$

Empirical Formula Pogil

R. Ayres



Empirical Formula Pogil:

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