

Empirical and Molecular Formulas Worksheet

Key

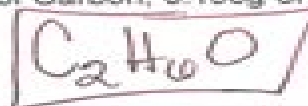
Objectives:

- be able to calculate empirical and molecular formulas

Empirical Formula

- 1) What is the empirical formula of a compound that contains 0.783g of Carbon, 0.196g of Hydrogen and 0.521g of Oxygen?

$$\begin{array}{l} \text{C} \cdot \frac{0.783\text{g C}}{12.01\text{g}} = \frac{0.065}{0.033} \approx 1.96 \quad 2 \\ \text{H} \cdot \frac{0.196\text{g H}}{1.01\text{g}} = \frac{0.194}{0.033} \approx 5.88 \quad 6 \\ \text{O} \cdot \frac{0.521\text{g O}}{16.00\text{g}} = \frac{0.033}{0.033} = 1 \end{array}$$



- 2) What is empirical formula of a compound which consists of 89.14% Au and 10.80% of O?

$$\text{Au} \frac{89.14\text{g}}{196.97} = \frac{0.45}{0.45} = 1 \quad \text{AuO}_{1.5}$$

$$\text{O} \frac{10.80\text{g}}{16\text{g}} = \frac{0.675}{0.45} = 1.5$$



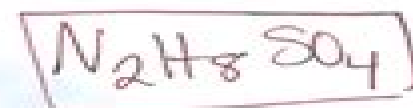
- 3) What is empirical formula if compound consists of 21.2%N, 6.1%H, 24.2%S and 48.5%O?

$$\text{N} \frac{21.2\text{g}}{14.01} = 1.5 \quad 2$$

$$\text{H} \frac{6.1\text{g}}{1.01\text{g}} = 6.03 \quad 8$$

$$\text{S} \frac{24.2\text{g}}{32} = 0.75 \quad 1$$

$$\text{O} \frac{48.5\text{g}}{16.00\text{g}} = 3.03 \quad 4$$



Molecular Formula

- 4) Empirical formula of a substance is CH_2O . Molar mass is 180. What is the molecular formula?

$$\text{EFM} = (12 + 2 \cdot 0 + 16) =$$

$$\frac{180}{30} = 6$$



- 5) Sample (3.585g) contains 1.388g of C, 0.345g of H, 1.850g O and its molar mass is 62g. What is molecular formula of this substance?

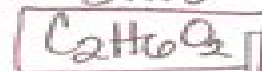
$$\text{C} \frac{1.388}{3.585} \rightarrow 38.7\% \rightarrow \frac{3.87\text{g}}{12.01\text{g}} = \frac{3.22}{3.22} = 1$$

$$\text{H} \frac{0.345}{3.585} \rightarrow \frac{9.62\text{g}}{1.01\text{g}} = \frac{9.52}{3.22} = 3$$

$$\text{O} \frac{1.850}{3.585} \rightarrow \frac{51.6\text{g}}{16\text{g}} = \frac{3.22}{3.22} = 1$$

$$\text{CH}_3\text{O} = \text{EFM} (12 + 3 \cdot 0 + 16)$$

$$\frac{62}{31.03} = 2$$



1	C ₂ H ₆ O
2	Na ₂ SO ₄
3	CH ₃ CO ₂ H
4	CH ₃ CO ₂ Na
5	CH ₃ CO ₂ CH ₃

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J Dewey



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