

# Chemistry Notes

## Molarity:

- Solution: a mixture of more than one component that can be separated by physical means such as filtration or distillation.
- Solvent : the compound of a solution

$$\text{Molarity (M): } \frac{\text{Mols}}{1\text{L}}$$

- Concentrated solutions: relatively large amount of solute
- Diluted solution: relatively small amount of solute

## Molarity calculations:

$$\text{A. } 0.500 \text{ mole in } 2.00\text{L solution} \quad \frac{0.500 \text{ moles}}{2.00 \text{ L solution}} = 0.25\text{L M}$$

$$\begin{aligned} \text{B. } 1.325 \text{ Moles in } 400.00 \text{ ml solution} \quad & \frac{400.00 \text{ ml}}{1000 \text{ ml}} = 0.400 \text{ M} \\ & \frac{1.325 \text{ mols}}{0.400 \text{ L}} = 3.31 \text{ M} \end{aligned}$$

# **Molarity And Dilution Review Document**

**Institute for Applied Technology  
(U.S.). Building Research  
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