Name Class

Assessment)

Chapter Test A

Chapter: Acids and Bases

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- Which of the properties listed below is not characteristic of an acid?
 - a. a sour taste
 - b. a slippery feel
 - c. the ability to conduct an electric current
 - d. reactivity with metals
 - 2. Which is an example of a weak acid?
 - a. HCI
 - **b.** H₂SO₄
 - c. HaPO₄
 - d. HNO_a
- _____ 3. Acetic acid is found in
 - a. vinegar.
 - b. the stomach.
 - c. antacids.
 - d. oranges.
- _____ 4. A characteristic of an Arrhenius base is that it
 - a. is an electrolyte.
 - b. tastes bitter.
 - c. dissociates to form OH ons.
 - d. All of the above
- 5. Which of the following is a Brønsted-Lowry base?
 - a. HCl
 - b. HCO_a
 - c. H₀O *
 - d. H_0PO_4

6. In the reaction represented by the equation

$$H_2C_2O_4(aq) + CH_3NH_2(aq) \Leftrightarrow HC_2O_4^-(aq) + CH_3NH_2^+(aq),$$

which of these is a conjugate acid-base pair?

- a. $H_2C_2O_4(aq)$ and $CH_3NH_2(aq)$
- b. H₂C₂O₄(aq) and CH₃NH^{*}₃(aq)
- c. CH₃NH₂(aq) and CH₃NH'₃(aq)
- d. HC2O4(aq) and CH2NH2(aq)

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the first step to writing a net ionic equation is balancing the chemical equation present let's use the reaction between sodium chloride and silver nitrate as an example next we write the

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