

A) oxidation. B) reduction. C) hydrolysis. D) neutralization. E) hydrogenation.

16. What must a monomer have to create a polymer by addition polymerization?
A) a dipole B) a cross-link C) a hydrogen bond D) a double carbon bond

17. In which one of the following reactions can Zaitsev's Rule be applied?
A) Hydrolysis of alkyl halides B) Hydration of but-2-ene
C) Dehydration of butan-2-ol D) Hydration of pentan-1-ene

18. Given the equation: This reaction is best described as:
 $C_6H_6 + Cl_2 \rightarrow C_6H_5Cl + HCl$
A) addition involving a saturated hydrocarbon
B) addition involving an unsaturated hydrocarbon
C) substitution involving a saturated hydrocarbon
D) substitution involving an unsaturated hydrocarbon

19. Which of the following is the electron configuration for magnesium?
A) $1s^2 2s^2 2p^6$ B) $1s^2 2s^2 2p^6 3s^2$ C) $1s^2 2s^2 2p^6 3s^1$
D) $1s^2 2s^2 2p^6 3s^1$ E) $1s^2 2s^2 2p^6 3s^2$

20. What is the maximum number of orbitals in a 5f subshell?
A) 2 B) 5 C) 7 D) 10 E) 14

21. Which one of the following is not a valid value for the magnetic quantum number of an electron in a 5d subshell?
A) 2 B) 3 C) 0 D) 1 E) -4

22. An electron cannot have the quantum numbers: $n = \underline{\hspace{1cm}}$, $l = \underline{\hspace{1cm}}$, $m_l = \underline{\hspace{1cm}}$.
A) 2, 0, 0 B) 2, 1, -1 C) 3, 1, -1 D) 1, 1, 1 E) 3, 2, 1

23. How many valence electrons does an element whose outer shell configuration is $3p^2$ have?
A) 4 B) 5 C) 6 D) 7

24. Which of the following correctly lists the five atoms in order of increasing size (smallest to largest)?
A) $O < F < S < Mg < Ba$ B) $F < O < S < Mg < Ba$
C) $F < O < S < Ba < Mg$ D) $O < F < S < Ba < Mg$
E) $F < S < O < Mg < Ba$

25. Identify the period and group numbers of the element with the electron configuration: $[Ne]3s^2 3p^2$

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Sabine Zange



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