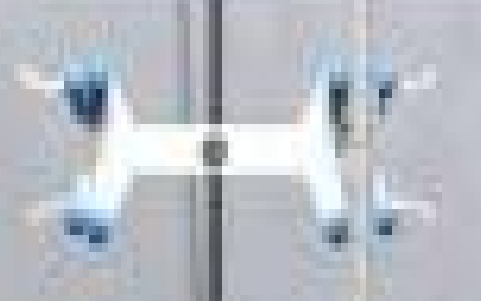


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By definition, the number of moles of reactant is equal to the volume (V) multiplied by the concentration (C):

$$n = C \times V$$

From the balanced chemical equation we know that the ratio of the two reacting compounds is 1 to 2:



So at the equivalence point:



Now you can calculate the unknown concentration of the H_2SO_4 according to:

$$C_{\text{H}_2\text{SO}_4} = \frac{1}{2} \times \frac{M_{\text{NaOH}} \times V_{\text{NaOH}}}{V_{\text{H}_2\text{SO}_4}}$$

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