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<sup>1</sup> Based on third-party testing from Quality Logic, commissioned by HP.

See full reports at <http://www.qualitylogic.com/Contents/Library/TestReports/HPInstantOnPerformanceReports.aspx>

<sup>2</sup> Results are based on internal HP testing using the ENERGY STAR® program's typical Electricity Consumption (TEC) method. Test data was extended to one year.

Testing was on a single unit of each product. Individual product configuration will affect power consumption.

<sup>3</sup> Available with dn and x models only.

<sup>4</sup> Not included, please purchase separately.

<sup>5</sup> The P3015d and dn models have a 600-sheet standard input capacity, two 500-sheet trays may be added, a model has an 1,100-sheet standard input capacity, an additional 500-sheet tray can be added.



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# Hp Laserjet 3015 Manual

**Ensheng Dong**



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web limiting reactant theoretical and percentage yields key  $4\text{KO}_2 + 2\text{H}_2\text{O} \rightarrow 4\text{KOH} + 3\text{O}_2$  g if a reaction vessel contains 0 mol  $\text{KO}_2$  and 0 mol  $\text{H}_2\text{O}$  what is the limiting reactant

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web start with 8.3 mole  $\text{H}_2\text{S}$  and multiply it with 2 mole  $\text{H}_2\text{O}$  over 2 mole  $\text{H}_2\text{S}$  from the equation  $8.3\text{ mole H}_2\text{S} \times \frac{2\text{ mole H}_2\text{O}}{2\text{ mole H}_2\text{S}} \times 18\text{g H}_2\text{O} = 149.4\text{g H}_2\text{O}$  is your theoretical yield so percent yield  $\frac{137.1\text{g H}_2\text{O}}{149.4\text{g H}_2\text{O}} \times 100 = 91.77\%$  problems 6.8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of  $\text{NH}_3$  can be produced theoretically from the reaction of 5.0 g of  $\text{N}_2$  what is the

limiting reagent if 8.52 g are actually formed what is the  
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web limiting reagents problem one 1 for the balanced equation  $\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$  if the reaction of 21.3 grams of  $\text{O}_2$  produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have  $\frac{21.3 \text{ g}}{32 \text{ g/mol}} = 0.666 \text{ mol}$

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web percent yield is a measure of the efficiency of the experimental design  $\text{yield} = \frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$  in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction  $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$  b determine the theoretical yield of  $\text{Al}_2\text{O}_3$  160.0 g Al  $\times \frac{1 \text{ mol Al}}{27 \text{ g Al}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{3 \text{ mol Al}} \times 101.96 \text{ g/mol} = 63.9 \text{ g}$  c determine the percent yield  $\frac{101.96 \text{ g}}{160.0 \text{ g}} \times 100 = 63.7\%$

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web 7 Na<sub>2</sub>C<sub>2</sub>O<sub>4</sub> is the limiting reactant percent yield 86.6% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4.33 g of H<sub>2</sub> are left over 13

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web the percent yield is calculated as follows  $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  based on this definition we would expect a percent yield

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reacted with 8.6 g of  $H_2$

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 $80.0 \text{ g } I_2O_5 \times \frac{1 \text{ mol } I_2O_5}{283.5 \text{ g } I_2O_5} \times \frac{1 \text{ mol } I_2}{1 \text{ mol } I_2O_5} \times \frac{253.8 \text{ g } I_2}{1 \text{ mol } I_2} = 70.9 \text{ g } I_2$

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 percent yield =  $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$

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