

Physical Science – 2nd Semester – Final Exam Study Guide – (2012–2013)

1. Physical Science Basics

- What tool(s) would you want to use to find the:
 - Mass of an object? triple beam balance Basic SI Unit? kg
 - Volume of an object? ruler or water displacement method Basic SI Unit? cm³ or ml
 - Length of an object? Ruler Basic SI Unit? meter
 - Density of an object? graduated cylinder Basic SI Unit? ml
- What is the formula for calculating density? mass/volume
 - What is the density of a cube that is 25 g and has a volume of 3 cm³? 8g/cm³
 - Compare the density of solid H₂O to liquid H₂O, and explain what observable effect this has.
solid H₂O less dense than liquid H₂O

2. Chemistry Basics

- Element (def.) substance made up of one kind of atom
- Atom (def.) basic building block of matter
- Compound (def.) a molecule that contains atoms or molecules of two or more different elements
- Molecule (def.) a combination of two or more of the same or different atoms that are bonded together

3. Model of the Atomic Evolution of the Atomic Theory

Atoms are so small that many scientists throughout history have created models to describe them.

Why have atomic theories changed over time? Better technology came around.

Scientist	Year	Major Discovery/ Contribution	Name of atomic model
John Dalton	1808	Each element is made up of small, indivisible particles called atoms. Different elements have atoms of different masses.	Billiard Ball Model
J.J. Thomson	1897	Negatively charged electron	Plum Pudding Model
Niels Bohr	1913	Electrons move <u>fixed orbits</u> around the atom's nucleus in specific layers, or <u>electron shells</u> . Atoms absorb or give off energy when the electrons move from one shell (<u>energy level</u>) to another.	Bohr Model
Ernest Rutherford	1909	1909 /Based on idea that positive reflects positive/ atom mostly empty space; electrons orbit randomly around positively charged nucleus	Planetary Model

4. Periodic Table

- The current periodic table is arranged by increasing atomic number.
- Reading the Periodic Table
 - An element's properties can be predicted from its location in the periodic table.
 - The atomic number represents the number of protons.
 - The atomic mass represents the number of protons and neutrons.

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Ryan**



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Pamphlets Including Serials and Contributions to Periodicals January June Chemical News and Journal of Physical
Science ,1903 *Chemical News and Journal of Physical Science* William Crookes, James H. Gardiner, Gerald Druce, H. W.
Blood-Ryan, 1919 **Course of Study in Physical Science for the Junior High School, Grades 7, 8, and 9** New Jersey.
Department of Public Instruction, 1927 *Teaching Reading in Secondary School Content Subjects* Carl Bernard
Smith, Sharon L. Smith, Larry Mikulecky, 1978 *The Publishers' Trade List Annual* ,1902 Resources for Teaching Middle
School Science Smithsonian Institution, National Academy of Engineering, National Science Resources Center of the National
Academy of Sciences, Institute of Medicine, 1998-03-30 With age appropriate inquiry centered curriculum materials and
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understanding of the world around them Resources for Teaching Middle School Science developed by the National Science
Resources Center NSRC is a valuable tool for identifying and selecting effective science curriculum materials that will
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National Science Education Standards This completely new guide follows on the success of Resources for Teaching
Elementary School Science the first in the NSRC series of annotated guides to hands on inquiry centered curriculum
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students can be expected to learn a list of accompanying materials a reading level and ordering information The curriculum
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guide The criteria reflect and incorporate goals and principles of the National Science Education Standards The annotations
designate the specific content standards on which these curriculum pieces focus In addition to the curriculum chapters the
guide contains six chapters of diverse resources that are directly relevant to middle school science Among these is a chapter
on educational software and multimedia programs chapters on books about science and teaching directories and guides to
science trade books and periodicals for teachers and students Another section features institutional resources One chapter

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 El-Hi Textbooks in Print ,1984 **The Science Teacher** ,1998 Some issues are accompanied by a CD ROM on a selected topic *Resources in Education* ,2000-10 Books in Print Supplement ,2002 *Textbooks in Print* ,1957
 Handbook on Contemporary Education Steven E. Goodman,1976 Science Curriculum Resource Handbook ,1992

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