

## CHAPTER 19 REVIEW

### *Oxidation-Reduction Reactions*

#### SECTION 1

**SHORT ANSWER** Answer the following questions in the space provided.

1. \_\_\_\_\_ All the following equations involve redox reactions *except*



2. Assign the correct oxidation number to the individual atom or ion below.

\_\_\_\_\_ a. Mn in  $\text{MnO}_2$

\_\_\_\_\_ b. S in  $\text{S}_8$

\_\_\_\_\_ c. Cl in  $\text{CaCl}_2$

\_\_\_\_\_ d. I in  $\text{IO}_3^-$

\_\_\_\_\_ e. C in  $\text{HCO}_3^-$

\_\_\_\_\_ f. Fe in  $\text{Fe}_2(\text{SO}_4)_3$

\_\_\_\_\_ g. S in  $\text{Fe}_2(\text{SO}_4)_3$

3. In each of the following half-reactions, determine the value of  $x$ .



\_\_\_\_\_ d. Which of the above half-reactions represent reduction processes?

4. Give examples, other than those listed in Table 1 on page 631 of the text, for the following:

\_\_\_\_\_ a. a compound containing H in a  $-1$  oxidation state

\_\_\_\_\_ b. a peroxide

\_\_\_\_\_ c. a polyatomic ion in which the oxidation number for S is  $+4$

\_\_\_\_\_ d. a substance in which the oxidation number for F is not  $-1$

# Holt Modern Chemistry Oxidation And Reduction

**L Cohen**



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