

Chemistry Worksheets Class 11 on Chapter 11 The p-Block Elements- Set 5

Q-1: Which of the following Group 14 elements has the lowest atomic volume?

- a) Si
- b) C
- c) Sn
- d) Ge

Q-2: Which of the following is an amorphous form of carbon?

- a) Graphite
- b) Lamp-black
- c) Diamond
- d) All of the above

Q-3: Which of the following types of graphite has an ABAB-type layer arrangement?

- a) α -graphite
- b) β -graphite
- c) Both a) and b)
- d) None of the above

Q-4: Conc. HNO_3 is stored in containers made up of

- a) Zinc
- b) Aluminium
- c) Tin
- d) Copper

Q-5: Which of the following is not a protonic acid?

- a) $\text{B}(\text{OH})_3$
- b) $\text{PO}(\text{OH})_3$
- c) $\text{SO}(\text{OH})_2$
- d) $\text{SO}_2(\text{OH})_2$

Q-6: Standard electrode potential values, E° for Al^{3+}/Al is -1.66 V and $+1.26 \text{ V}$ for Ti^{2+}/Ti . Predict the creation of the M^{2+} ion in solution and contrast the electropositive properties of the two metals.

Q-7: Boron is unable to form BF_6^{3-} ion. Explain.

Q-8: Why is boric acid regarded as a weak acid?

Episode 11chemistry Worksheet

Rachel S Tattersall



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