

# I. The Reaction Quotient

1. We can write K only for a reaction at equilibrium
2. How can we describe a reaction at other times?  $\text{N}_2 + 3\text{H}_2 \rightleftharpoons 2\text{NH}_3$

$$Q = \frac{[\text{NH}_3]_0^2}{[\text{N}_2]_0[\text{H}_2]_0^3} \quad \text{Describes the reaction when not at equilibrium}$$

3. **Q = reaction quotient** has the same form as K, but is at non-equilibrium times.
4. Q can tell us how a reaction will change to get to equilibrium
  - 1) If  $Q = K$ , we are at equilibrium
  - 2) If  $Q > K$ , then products > reactants and the reaction will shift left
  - 3) If  $Q < K$ , then reactants > products and the reaction will shift right

Q	K	Predicted Direction of Reaction
0.55	1.45	To the right (towards products)
2.55	1.45	To the left (towards reactants)
1.45	1.45	No change (at equilibrium)

5. *Example:* For the above reaction,  $K = 6 \times 10^{-2}$ . Predict change:
  - 1)  $[\text{N}_2]_0 = 1 \times 10^{-5} \text{ M}$ ,  $[\text{H}_2]_0 = 0.002 \text{ M}$ ,  $[\text{NH}_3]_0 = 0.001 \text{ M}$
  - 2)  $[\text{N}_2]_0 = 1.5 \times 10^{-5} \text{ M}$ ,  $[\text{H}_2]_0 = 0.354 \text{ M}$ ,  $[\text{NH}_3]_0 = 0.0002 \text{ M}$
  - 3)  $[\text{N}_2]_0 = 5 \text{ M}$ ,  $[\text{H}_2]_0 = 0.01 \text{ M}$ ,  $[\text{NH}_3]_0 = 0.0001 \text{ M}$

# Entropy Pogil Answer Key

**LP Steffe**



## Entropy Pogil Answer Key:

*Analytical Chemistry* Juliette Lantz, Renée Cole, The POGIL Project, 2014-12-31 An essential guide to inquiry approach instrumental analysis Analytical Chemistry offers an essential guide to inquiry approach instrumental analysis collection The book focuses on more in depth coverage and information about an inquiry approach This authoritative guide reviews the basic principles and techniques Topics covered include method of standard the microscopic view of electrochemistry calculating cell potentials the BerriLambert atomic and molecular absorption processes vibrational modes mass spectra interpretation and much more

*Entropy Demystified* Arieh Ben-Naim, 2007 In this unique book Arieh Ben Naim invites the reader to experience the joy of appreciating something which has eluded understanding for many years OCo entropy and the Second Law of Thermodynamics The book has a two pronged message first that the Second Law is not OC infinitely incomprehensibleOCO as commonly stated in textbooks of thermodynamics but can in fact be comprehended through sheer common sense and second that entropy is not a mysterious quantity that has OC resisted understandingOCO but a simple familiar and easily comprehensible concept Written in an accessible style the book guides the reader through an abundance of dice games and examples from everyday life The author paves the way for readers to discover for themselves what entropy is how it changes and most importantly why it always changes in one direction in a spontaneous process

Discover Entropy and the Second Law of Thermodynamics Arieh Ben-Naim, 2010 This is a sequel to the author s book entitled *Entropy Demystified* The aim is essentially the same as that of the previous book by the author to present Entropy and the Second Law as simple meaningful and comprehensible concepts In addition this book presents a series of experiments which are designed to help the reader discover entropy and the Second Law While doing the experiments the reader will encounter no unexpected results and concepts of entropy and the Second Law will emerge naturally from these experiments without a tinge of mystery These concepts are explained with the help of a few familiar ideas of probability and a 20 question game The main value of the book is to introduce entropy and the Second Law in simple language which renders it accessible to any reader who can read and is curious about the basic laws of nature The book is addressed to anyone interested in science and in understanding natural phenomenon It will give the reader the opportunity to discover one of the most fundamental laws of physics a law that has resisted complete understanding for over a century The book is also designed to be enjoyable There is no other book of its kind except *Entropy Demystified* by the same author that offers the reader a unique opportunity to discover one of the most profound laws sometimes viewed as a mysterious law without the tinge mystery There are no pre requisites expected of the readers all that the reader is expected to do is to follow the experiments or imagine doing the experiments and reach the inevitable conclusions

Entropy Demystified: The Second Law Reduced To Plain Common Sense (Revised Edition) Arieh Ben-naim, 2008-06-18 In this unique book the reader is invited to experience the joy of appreciating something which has eluded understanding for many years entropy and the Second Law of Thermodynamics

The book has a two pronged message first that the second law is not infinitely incomprehensible as commonly stated in most textbooks on thermodynamics but can in fact be comprehended through sheer common sense and second that entropy is not a mysterious quantity that has resisted understanding but a simple familiar and easily comprehensible concept Written in an accessible style the book guides the reader through an abundance of dice games and examples from everyday life The author paves the way for readers to discover for themselves what entropy is how it changes and most importantly why it always changes in one direction in a spontaneous process In this new edition seven simulated games are included so that the reader can actually experiment with the games described in the book These simulated games are meant to enhance the readers understanding and sense of joy upon discovering the Second Law of Thermodynamics Entropy And Its Physical Meaning

J. S. Dugdale,1996-08-28 This text gives students a clear and easily understood introduction to entropy a central concept in thermodynamics but one which is often regarded as the most difficult to grasp Professor Dugdale first presents a classical and historical view of entropy looking in detail at the scientists who developed the concept and at how they arrived at their ideas This is followed by a statistical treatment which provides a more physical portrait of entropy relating it to disorder and showing how physical and chemical systems tend to states of order at low temperatures Dugdale includes here a brief account of some of the more intriguing manifestations of order in properties such as superconductivity and superfluidity Entropy and Its Physical Meaning also includes a number of exercises which can be used for both self learning and class work It is intended to provide a complete understanding of the concept of entropy making it valuable reading for undergraduates in physics physical sciences and engineering and for students studying thermodynamics within other science courses such as meteorology biology and medicine **Entropy and the Second Law** Arieh Ben-Naim,2012-05-15 This book

presents a clear and readable description of one of the most mysterious concepts of physics Entropy It contains a self learning kit that guides the reader in understanding the concepts of entropy In the first part the reader is asked to play the familiar twenty Question game Once the reader feels comfortable with playing this game and acquires proficiency in playing the game effectively intelligently he or she will be able to capture the elusive and used to be mysterious concept of entropy There will be no more speculative or arbitrary interpretations nor older or modern views of entropy This book will guide readers in choosing their own interpretation of entropy Video intro on the Bestsellers on Entropy by Arieh Ben Naim <https://www.youtube.com/watch?v=S5fOsKyOIHW> Request Inspection Copy Contents Introduction From Heat Engines to Disorder Information Spreading Freedom and More Forget about Entropy for a While Let us Go and Play iGamesThe Astounding Emergence of the Entropy of a Classical Ideal Gas out of Shannon's Measure of InformationExamples and Their Interpretations Challenges for any Descriptor of EntropyFinally Let Us Discuss the Most Mysterious Second Law of Thermodynamics Readership Undergraduate and graduate students in chemistry and physics academics and lay persons

**A Student's Guide to Entropy** Don Stephen Lemons,2013-08-29 This book helps readers understand the elusive

concept of entropy to supplement undergraduate courses in physics engineering chemistry and mathematics *Entropy and Energy Levels* Robert Paul Holland Gasser, William Graham Richards, 1974      **Introduction to Entropy** Jonathan Allday, Simon Hands, 2024-10-07 The concept of entropy arises in diverse branches of science including physics where it plays a crucial role However the nature of entropy as a unifying concept is not widely discussed it is dealt with in a piecemeal manner within different contexts The interpretation of the concept is also subtly different in each case This book will draw these diverse threads together and present entropy as one of the crucial physical concepts It will cover a range of different applications of entropy from the classical theory of thermodynamics the statistical approach entropy in quantum theory information theory and finally its manifestation in black hole physics Each will be presented in a manner suitable for undergraduates and interested laypersons with no previous knowledge The book will take an overview of these areas and see to what extent the concept of entropy is being treated in the same way in each and how it differs Key Features Provides an accessible introduction to the exciting topic of entropy setting out its manifestations in classical thermodynamics statistical mechanics and information theory Covers applications in black holes quantum theory and Big Bang cosmology      **Entropy Analysis** Norman C. Craig, 1997-06      *Entropy, Specific Heat, Etc., Data on U, UO<sub>2</sub>, and UO<sub>3</sub>* K. K. Kelley, 1947      Entropy Johan Diedrich Fast, 1968      *Entropy Analysis* Norman C. Craig, 1992      What is Entropy? Entropy Press, 1995      **The Entropy Diagram and Its Applications** Jules Boulvin, 1914      **Concept Learning and Problem Solving** John Francis Cullen, 1983      *Maximum Entropy Solutions to Scientific Problems* R. M. Bevensee, 1993 Highlighting several versions of the flexible maximum entropy ME method this reference provides strategies for solving various practical inverse and undetermined problems It explores the advantages and disadvantages of using different methods and backs up solutions with specific examples      Entropy Lifeliqe, 2019 This lesson plan covers entropy one of the forces that drive chemical reactions

**Entropy : the Significance of the Concept of Entropy and Its Applications in Science and Technology** J.D. Fast, 1968      **The Method of Maximum Entropy** Henryk Gzyl, 1995 This monograph is an outgrowth of a set of lecture notes on the maximum entropy method delivered at the 1st Venezuelan School of Mathematics This yearly event aims at acquainting graduate students and university teachers with the trends techniques and open problems of current interest In this book the author reviews several versions of the maximum entropy method and makes its underlying philosophy clear

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