

1. Which electron configuration represents an atom in an excited state?

- A)  $1s^2 2s^2 2p^6 3p^1$  C)  $1s^2 2s^2 2p^6 3s^2 3p^2$   
B)  $1s^2 2s^2 2p^6 3s^2 3p^1$  D)  $1s^2 2s^2 2p^6 3s^2$

2. All of the elements in Period 3 have a total of 2 electrons in the

- A) 2s sublevel C) 2p sublevel  
B) 3s sublevel D) 3p sublevel

3. Which atom in the ground state has three unpaired electrons in its outermost principal energy level?

- A) Li C) N  
B) B D) Ne

4. What is the total number of valence electrons in an atom with the electron configuration  $1s^2 2s^2 2p^6 3s^2 3p^3$ ?

- A) 6 C) 3  
B) 2 D) 5

5. Which orbital notation represents an atom of beryllium in the ground state?

- A) 

1s	2s	2p		
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- B) 

1s	2s	2p		
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- C) 

1s	2s	2p		
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- D) 

1s	2s	2p		
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6. Which element has a completely filled third principal energy level?

- A) Ar C) Fe  
B) N D) Zn

7. Which atom in the ground state has only 3 electrons in the 3p sublevel?

- A) phosphorus C) argon  
B) potassium D) aluminum

8. What is the total number of occupied principal energy levels in a neutral atom of neon in the ground state?

- A) 1 C) 3  
B) 2 D) 4

9. Which sublevel configuration correctly represents a completely filled third principal energy level?

- A)  $3s^2 3p^6 3d^6$  C)  $3s^2 3p^6 3d^{10}$   
B)  $3s^2 3p^2 3d^{10}$  D)  $3s^2 3p^6 3d^6$

10. Which atom in the ground state contains a partially filled 3p orbital?

- A) argon C) potassium  
B) calcium D) aluminum

11. In the ground state, the atoms of elements in Period 2 all have the same number of

- A) protons C) 1s electrons  
B) neutrons D) oxidation states

12. Which is the electron configuration for a neutral atom in the ground state?

- A)  $1s^2 2s^2 3s^1$  C)  $1s^2 2s^2 2p^6 3p^1$   
B)  $1s^2 2s^2 2p^4 3s^1$  D)  $1s^2 2s^2 2p^6 3s^1$

13. Which represents the electron configuration of an isotope of oxygen in the ground state?

- A)  $1s^2 2s^2 2p^1$  C)  $1s^2 2s^2 2p^2$   
B)  $1s^2 2s^2 2p^2$  D)  $1s^2 2s^2 2p^4$

14. The total number of sublevels in the fourth principal energy level of an atom is

- A) 1 C) 3  
B) 2 D) 4

15. Which is a possible electron configuration for argon in the excited state?

- A)  $1s^2 2s^2 2p^7 3s^2 3p^2$  C)  $1s^2 2s^2 2p^6 3s^1 3p^2$   
B)  $1s^2 2s^2 2p^5 3s^2 3p^6$  D)  $1s^2 2s^2 2p^6 3s^2 3p^5 4s^1$

16. Which is a correct description of the shape and spatial orientation of the p orbitals in an atom?

- A) All have the same shape but a different orientation.  
B) All have the same shape and the same orientation.  
C) All have a different shape and a different orientation.  
D) All have a different shape but the same orientation.

17. A maximum of 6 electrons can occupy

- A) an s orbital C) a p orbital  
B) an s sublevel D) a p sublevel

18. How does the ground state electron configuration of the hydrogen atom differ from that of a ground state helium atom?

- A) Hydrogen has one electron in a higher energy level.  
B) Hydrogen has two electrons in a lower energy level.  
C) Hydrogen contains a half-filled orbital.  
D) Hydrogen contains a completely filled orbital.

# Electron Distribution Review Worksheet Answers

**Tobias Bleicker**



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