

A silver and red DC Metro bus is parked on a city street. The bus has "metro" written on its side in blue. The background shows a multi-story brick building with many windows. The bus is facing left, and its front end is visible. The text "RIDING THE DC METRO BUS: A BEGINNER'S GUIDE" is overlaid in the center of the image.

RIDING THE DC METRO BUS: A BEGINNER'S GUIDE

Metro Bus Ride Guide

N Noddings



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$$\begin{array}{l} 1 \text{ N}_2 + \text{F}_2 \rightarrow 2 \text{ NF}_3 \\ 2 \text{ C}_6\text{H}_{10} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O} \\ 3 \text{ HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2 \\ \text{GaBr}_3 + \text{NaBr} \rightarrow \text{SNO}_4 + \text{NaSO}_3 + \text{SO}_3 \\ 3 \text{ NF}_3 + \text{SNF}_2 \rightarrow \text{N}_2\text{O}_3 \end{array}$$
 using the equation from problem 2 above answer the following

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$$\begin{array}{l} 1 \text{ N}_2 + 3 \text{ F}_2 \rightarrow 2 \text{ NF}_3 \\ 2 \text{ C}_6\text{H}_{10} + 17 \text{ O}_2 \rightarrow 12 \text{ CO}_2 + 10 \text{ H}_2\text{O} \\ 3 \text{ HBr} + 1 \text{ KHCO}_3 \rightarrow 1 \text{ H}_2\text{O} + 1 \text{ KBr} + \text{CO}_2 \\ 2 \text{ GaBr}_3 + 3 \text{ Na}_2\text{SO}_3 \rightarrow 1 \text{ Ga}_2\text{SO}_3 + 3 \text{ SO}_3 + 6 \text{ NaBr} + 3 \text{ SNO}_2 + 3 \text{ NF}_3 + 3 \text{ SNF}_2 + 1 \text{ N}_2\text{O}_3 \end{array}$$
 using the following equation $2 \text{ NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2 \text{ H}_2\text{O} + \text{Na}_2\text{SO}_4$

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$$\text{Ce} + \text{Zn} + \text{CuCl}_2 \rightarrow \text{ZnCl}_2 + \text{Cu} + \text{Zn}$$

 how many moles of Ce will be produced from 23.0 g of Ce assuming CuCl_2 is available in excess

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$$\text{N}_2 + 3\text{F}_2 \rightarrow 2\text{NF}_3$$

$$2\text{C}_6\text{H}_{10} + 17\text{O}_2 \rightarrow 12\text{CO}_2 + 10\text{H}_2\text{O}$$

$$3\text{H}_2\text{O} + 1\text{KBr} + \text{CO}_2 \rightarrow 4\text{H}_2\text{O} + 2\text{GaBr}_3 + 3\text{Na}_2\text{SO}_3 + 1\text{Ga}_2\text{SO}_3 + 3\text{NaBr}$$

$$5\text{S} + 3\text{NO}_2 + 3\text{NF}_3 \rightarrow 2\text{N}_2\text{O}_3 + 3\text{SnF}_2 + 1\text{N}_2\text{O}_3$$

 using the following equation

$$\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{Na}_2\text{SO}_4$$

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$$\text{N}_2 + 3\text{F}_2 \rightarrow 2\text{NF}_3$$

$$2\text{C}_6\text{H}_{10} + 17\text{O}_2 \rightarrow 12\text{CO}_2 + 10\text{H}_2\text{O}$$

$$\text{H}_2\text{O} + \text{KBr} + \text{CO}_2 \rightarrow 4\text{H}_2\text{O} + 2\text{GaBr}_3 + 3\text{Na}_2\text{SO}_3 + 1\text{Ga}_2\text{SO}_3 + 3\text{NaBr}$$

$$5\text{S} + 3\text{NO}_2 + 3\text{NF}_3 \rightarrow 2\text{N}_2\text{O}_3 + 3\text{SnF}_2 + 1\text{N}_2\text{O}_3$$

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$$\text{SiO}_2 + 3\text{C} \rightarrow \text{SiC} + 2\text{CO}$$

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$$\text{N}_2 + 3\text{F}_2 \rightarrow 2\text{NF}_3$$

$$2\text{C}_6\text{H}_{10} + 17\text{O}_2 \rightarrow 12\text{CO}_2 + 10\text{H}_2\text{O}$$

$$\text{H}_2\text{O} + \text{KBr} + \text{CO}_2 \rightarrow 4\text{H}_2\text{O} + 2\text{GaBr}_3 + 3\text{Na}_2\text{SO}_3 + 1\text{Ga}_2\text{SO}_3 + 3\text{NaBr}$$

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