

Identification of a Metal Carbonate

Experiment 3: A study in stoichiometry

Metal carbonates react with hydrochloric acid to produce water and carbon dioxide gas. This is the "acid test" used by geologists to identify carbonate containing rocks and minerals. The equation for this reaction, as it pertains to this experiment, is:



or for a bicarbonate (hydrogen carbonate):



Carbon dioxide is fairly soluble in water under high pressure conditions, as evidenced in carbonated soft drinks, however at normal atmospheric pressure the carbon dioxide escapes and the carbonated beverage goes flat if left without being tightly capped.

In this experiment, a weighted sample of unknown metal carbonate will chemically react and the mass of carbon dioxide escaping will be determined. This will provide the data needed to calculate the molar mass of the unknown and determine which of the substances listed in Table 1 can be identified as the unknown metal carbonate. **WARNING!** Some of the unknowns as well as the hydrochloric acid used in this experiment are very corrosive substances. Wear Eye Protection and avoid skin contact. Wash with running water if exposed.

PRE-LAB ASSIGNMENT: In the space provided in Table 1, write the chemical name and molar mass for each of the following carbonates.

Table 1: Metal Carbonate Unknowns

CHEMICAL NAME	CHEMICAL FORMULA	MOLAR MASS
	Na₂CO₃	
	Na₂CO₃ · ½ H₂O	
	NaHCO₃	
	K₂CO₃	
	KHCO₃	
	K₂CO₃ · ½ H₂O	

Metal Carbonate With Hydrochloric Acid Experiment Isa

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