

Chapter 8 Chemical Reactions Chapter Review Answers

1. The evolution of energy, the production of a gas, the formation of a precipitate, and a color change.
2. A correctly written chemical equation must represent the known facts, contain the correct formulas for the reactants and products, and satisfy the law of conservation of mass.
3.
 - a. It is a number that appears in front of a formula in a chemical equation.
 - b. The number of atoms of each type in the formula is multiplied by the coefficient.
4. methane + oxygen $\rightarrow$ carbon dioxide + water;
 $\text{CH}_4(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g})$;
 $\text{CH}_4(\text{g}) + 2\text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g})$
5. A chemical equation provides the number of molecules, moles, and atoms of each element or compound composing the reactants and products as well as the mole ratios of the reactants to each other and to the products.
6. Word and formula equations have only qualitative meaning; they do not express the relative quantities of a reaction's reactants or products.
7.
 - a. a solution in water
 - b. a substance that accelerates a chemical reaction but can be recovered unchanged
 - c. a chemical reaction in which the products re-form the original reactants
8.
 - a. KOH
 - b. $\text{Ca}(\text{NO}_3)_2$
 - c. Na_2CO_3
 - d. CCl_4
 - e. MgBr_2
9. (1) Balance different types of atoms one at a time; (2) balance types of atoms that appear only once on each side of the equation first; (3) balance as single units any polyatomic ions that appear on both sides of the equation; and (4) balance H atoms and O atoms last.
10.
 - a. 6N
 - b. 4H, 2O
 - c. 4H, 4N, 12O
 - d. 2Ca, 4O, 4H
 - e. 3Ba, 6Cl, 18O
 - f. 5Fe, 10N, 30O
 - g. 12Mg, 8P, 32O
 - h. 4N, 16H, 2S, 8O
 - i. 12Al, 188Se, 72O
 - j. 12C, 32H
11.
 - a. $2\text{ZnS}(\text{s}) + 3\text{O}_2(\text{g}) \rightarrow 2\text{ZnO}(\text{s}) + 2\text{SO}_2(\text{g})$
 - b. $2\text{HCl}(\text{aq}) + \text{Ba}(\text{OH})_2(\text{aq}) \rightarrow \text{BaCl}_2(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$
 - c. $2\text{HNO}_3(\text{aq}) + \text{Ca}(\text{OH})_2(\text{aq}) \rightarrow \text{Ca}(\text{NO}_3)_2(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$
12.
 - a. Solid zinc sulfide reacts with oxygen gas to form solid zinc oxide and sulfur dioxide gas.
 - b. When solid calcium hydride is added to water, aqueous calcium hydroxide and hydrogen gas are formed.
 - c. Aqueous silver nitrate mixed with aqueous potassium iodide produces a precipitate of silver iodide and aqueous potassium nitrate.
13.
 - a. $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$
 - b. $2\text{Al} + \text{Fe}_2\text{O}_3 \rightarrow \text{Al}_2\text{O}_3 + 2\text{Fe}$
 - c. $\text{Pb}(\text{CH}_3\text{COO})_2 + \text{H}_2\text{S} \rightarrow \text{PbS} + 2\text{CH}_3\text{COOH}$
14.
 - a. LiO_2 is an incorrect formula; $4\text{Li} + \text{O}_2 \rightarrow 2\text{Li}_2\text{O}$
 - b. H_2Cl_2 is an incorrect formula; $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$
 - c. MgO_2 is an incorrect formula, and the equation as written is not balanced; $\text{MgCO}_3 \rightarrow \text{MgO} + \text{CO}_2$
 - d. I is an incorrect formula for iodine, and the equation is not balanced; $2\text{NaI} + \text{Cl}_2 \rightarrow 2\text{NaCl} + \text{I}_2$
15.
 - a. $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$
 - b. $\text{P}_2\text{O}_5 + 6\text{H}_2\text{O} \rightarrow 4\text{H}_3\text{PO}_4$
 - c. $\text{Fe}_2\text{O}_3 + 3\text{CO} \rightarrow 2\text{Fe} + 3\text{CO}_2$
16. $\text{CH}_4(\text{g}) + 4\text{Cl}_2(\text{g}) \rightarrow \text{CCl}_4(\text{l}) + 4\text{HCl}(\text{g})$
17.
 - a. $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$
 - b. $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$
 - c. $2\text{Li} + \text{Cl}_2 \rightarrow 2\text{LiCl}$
 - d. $\text{Ca} + \text{I}_2 \rightarrow \text{CaI}_2$
18.
 - (1) In a synthesis reaction, two or more substances combine to form a new substance: $\text{A} + \text{X} \rightarrow \text{AX}$
 - (2) In a decomposition reaction, a single compound undergoes a reaction that produces two or more simpler substances: $\text{AX} \rightarrow \text{A} + \text{X}$
 - (3) In a single-displacement reaction, one element replaces another element in a compound:
 $\text{A} + \text{BX} \rightarrow \text{AX} + \text{B}$ or $\text{Y} + \text{RX} \rightarrow \text{RY} + \text{X}$

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Lei Huang



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Table of Contents **Merrill Chemistry Ch 8 Review Answers**

1. Understanding the eBook **Merrill Chemistry Ch 8 Review Answers**
 - The Rise of Digital Reading **Merrill Chemistry Ch 8 Review Answers**
 - Advantages of eBooks Over Traditional Books
2. Identifying **Merrill Chemistry Ch 8 Review Answers**
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an **Merrill Chemistry Ch 8 Review Answers**
 - User-Friendly Interface
4. Exploring eBook Recommendations from **Merrill Chemistry Ch 8 Review Answers**
 - Personalized Recommendations
 - **Merrill Chemistry Ch 8 Review Answers** User Reviews and Ratings
 - **Merrill Chemistry Ch 8 Review Answers** and Bestseller Lists

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 - Merrill Chemistry Ch 8 Review Answers Budget-Friendly Options
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 - Merrill Chemistry Ch 8 Review Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
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 - Highlighting and Note-Taking Merrill Chemistry Ch 8 Review Answers
 - Interactive Elements Merrill Chemistry Ch 8 Review Answers
8. Staying Engaged with Merrill Chemistry Ch 8 Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Merrill Chemistry Ch 8 Review Answers
9. Balancing eBooks and Physical Books Merrill Chemistry Ch 8 Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Merrill Chemistry Ch 8 Review Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Merrill Chemistry Ch 8 Review Answers
 - Setting Reading Goals Merrill Chemistry Ch 8 Review Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Merrill Chemistry Ch 8 Review Answers
 - Fact-Checking eBook Content of Merrill Chemistry Ch 8 Review Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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web evidence of the struggle between the danes and germans from the 9th to the 12th century lies west of the town of schleswig here the danish kings built an impressive fortification wall known as the danewerk nearby are the ruins of haithabu a historic viking trading settlement evidence of both german and danish cultural influences abound throughout

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schleswig and holstein were caught between the rising nationalism and desire for unification of the prussian german nation
states and the traditional alliances with the danish kingdom the schleswig holsteiners