

II B.Tech I Semester Examinations, MAY 2011

CHEMICAL PROCESS CALCULATIONS

Chemical Engineering

Time: 3 hours

Max Marks: 75

Answer any FIVE Questions
All Questions carry equal marks

- (a) A gas mixture contains 0.1234 kg-mole of HCl, 0.153 Kg-mole of nitrogen, and 0.04037 Kg-mole of oxygen. Calculate its density in kg/m^3 at a pressure of 1.72 atm and a temperature of 40°C .

(b) State and prove law of Amagat. [7+8]
- (a) By titration, it was found that a sample of water contains alkalinity equivalent to 800 mg/lit CaCO_3 . Assuming that the water contains temporary hardness due to $\text{Mg}(\text{HCO}_3)_2$, find the $\text{Mg}(\text{HCO}_3)_2$ content of water.

(b) Explain Baume and API. [15]
- Pure sulfur is burned in a burner at the rate of 0.3 kg/s. fresh air is supplied at 303K and 100 kpa. The gasses from the burner contain 16.5% SO_2 , 3% O_2 and rest N_2 on SO_2 free volume basis. The gasses leave the burner at 1073k and 101.325 kpa. Calculate

 - The fraction of sulfur burned in SO_3
 - The % excess air over the amount required to oxidize the sulfur to SO_2
 - The volume of dry air in m^3/s (d). The volume burner gas in m^3/s . [15]
- (a) Discuss in detail what is heat of formation and how to calculate heat of formation.

(b) Calculate heat of reaction of the following
 $\text{HCl}(\text{g}) + \text{NH}_3(\text{g}) \rightarrow \text{NH}_4\text{Cl}(\text{s})$
 $\text{HCl}(\text{g})$ of change of enthalpy is -22,063 cal
 $\text{NH}_3(\text{g})$ of change of enthalpy is -11040 cal
 $\text{NH}_4\text{Cl}(\text{s})$ of change of enthalpy is -75,380 cal. [15]
- (a) What do you mean by crystallization? Explain how it is used in industries.

(b) What do you mean by dissolution. [15]
- Calculate the number of kilocalories required to heat, from 500 to 1500°C , 1 Cubic meter (standard conditions) of a gas having the following composition by volume :

[15]

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Thomas Griffiths



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