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SL Paper 2

A student determined the percentage of the active ingredient magnesium hydroxide, $\text{Mg}(\text{OH})_2$, in a 1.24 g antacid tablet.

The antacid tablet was added to 50.00 cm^3 of $0.100 \text{ mol dm}^{-3}$ sulfuric acid, which was in excess.

- Calculate the amount, in mol, of H_2SO_4 . [1]
- Formulate the equation for the reaction of H_2SO_4 with $\text{Mg}(\text{OH})_2$. [1]
- The excess sulfuric acid required 20.80 cm^3 of $0.1133 \text{ mol dm}^{-3}$ NaOH for neutralization.
Calculate the amount of excess acid present. [1]
- Calculate the amount of H_2SO_4 that reacted with $\text{Mg}(\text{OH})_2$. [1]
- Determine the mass of $\text{Mg}(\text{OH})_2$ in the antacid tablet. [2]
- Calculate the percentage by mass of magnesium hydroxide in the 1.24 g antacid tablet to three significant figures. [1]

Markscheme

- a. $n(\text{H}_2\text{SO}_4) = 0.0500 \text{ dm}^3 \times 0.100 \text{ mol dm}^{-3} = 0.00500/5.00 \times 10^{-3} \text{ mole}$

[1 mark]

- b. $\text{H}_2\text{SO}_4(\text{aq}) + \text{Mg}(\text{OH})_2(\text{s}) \rightarrow \text{MgSO}_4(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$

Accept an ionic equation.

[1 mark]

- c. $n(\text{H}_2\text{SO}_4) = \frac{1}{2} \times n(\text{NaOH}) = \frac{1}{2} (0.02080 \text{ dm}^3 \times 0.1133 \text{ mol dm}^{-3}) =$

$0.001178/1.178 \times 10^{-3} \text{ mole}$

[1 mark]

- d. $n(\text{H}_2\text{SO}_4) \text{ reacted} = 0.00500 - 0.001178 = 0.00382/3.82 \times 10^{-3} \text{ mole}$

[1 mark]

- e. $n(\text{Mg}(\text{OH})_2) = n(\text{H}_2\text{SO}_4) = 0.00382/3.82 \times 10^{-3} \text{ mole}$

$m(\text{Mg}(\text{OH})_2) = 0.00382 \text{ mol} \times 58.33 \text{ g mol}^{-1} = 0.223 \text{ g}$

Ib Chemistry 2sl Past Paper 1

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A red circular graphic with a gradient, appearing as a partial circle or a thick curved line, located to the right of the text box.

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