

Question Number	Acceptable Answers	Reject	Mark
22(b)(iii)	$\Delta H_{\text{reaction}} = + \frac{2.3(10)}{0.2(00)} \times 2 = +23.1 \text{ (kJ mol}^{-1}\text{)}$ $= +23 \text{ (kJ mol}^{-1}\text{) to 2 sf}$ <p>First mark – correct computation of $\Delta H_{\text{reaction}}$:</p> <p>2 x [answer to (b)(i) in kJ ÷ answer to (b)(ii) in mol] (1)</p> <p>Second mark – stand alone, for correct rounding:</p> <p>A final answer to two sf (1)</p> <p>Third mark – stand alone, for giving a + sign for endothermic reaction:</p> <p>+ sign in front of final answer (1)</p> <p>NOTE: + 12 (kJ mol⁻¹) scores (2) (i.e. the 2nd and 3rd marks)</p>	Incorrect units given by the candidate (no 3 rd scoring point)	3

Question Number	Acceptable Answers	Reject	Mark
22(c)(i)	<p>(Average amount of) energy/enthalpy required to break one mole of (covalent) bonds</p> <p>ALLOW Energy change/enthalpy change to break one mole of (covalent) bonds (1)</p> <p>(in the) gas / gaseous (state) (1)</p>	<p>Energy/enthalpy released OR 'Bonds formed/made' OR 1 mol of compound for 1st mark</p>	2

Ial Edexcel Jan 2015 Papers

Minjie Lin



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- Chapter 2: Essential Elements of Ial Edexcel Jan 2015 Papers
- Chapter 3: Ial Edexcel Jan 2015 Papers in Everyday Life
- Chapter 4: Ial Edexcel Jan 2015 Papers in Specific Contexts
- Chapter 5: Conclusion

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