



GREEN Multiplying

$$1) -2(4x^2 - 5x) - x(x^2 + 6x)$$

$$-8x^2 + 10x - x^3 - 6x^2$$

$$-x^3 - 14x^2 + 10x$$

$$2) (w+5)(w^2+4w-8)$$

$$w^3 + 4w^2 - 8w + 5w^2 + 20w - 40$$

$$w^3 + 9w^2 + 12w - 40$$

$$3) (a-7)(a-7)$$

$$a^2 - 7a - 7a + 49$$

$$a^2 - 14a + 49$$

$$4) (x+3)(x+9)$$

$$x^2 + 9x + 3x + 27$$

$$x^2 + 12x + 27$$

$$5) (5m+3)(m-2)$$

$$5m^2 - 10m + 3m - 6$$

$$5m^2 - 7m - 6$$



BLUE Dividing

$$1) \frac{12c^5d^4 + 18c^4d^3}{3c^2d^3}$$

$$4c^3d + 6c^2$$

$$2) \frac{2x^3 - 18x^2}{2x^2} =$$

$$x - 9$$

$$3) \frac{4x^3 - 3x^2 + x}{x}$$

$$4x^2 - 3x + 1$$

$$4) \frac{16x^6 - 28x^4 + 8x^2}{4x^2}$$

$$4x^4 - 7x^2 + 2$$

$$5) \frac{27x^4 - 9x^3 + 63x^2}{9x}$$

$$3x^3 - x^2 + 7x$$

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RS Peters



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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_2(\text{g}) + \frac{5}{2} \text{O}_2(\text{g}) \quad \Delta h = 1300 \text{ kJ}$
 $\text{C}_2\text{H}_2(\text{g}) + 2 \text{ H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g}) \quad \Delta h = -84.5 \text{ kJ}$
 $\text{H}_2\text{O}(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l}) \quad \Delta h = -285 \text{ kJ}$
 $\text{CO}_2(\text{g}) + 3 \text{ H}_2(\text{g}) + \text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g}) \quad \Delta h = 283 \text{ kJ}$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 2 \text{ H}_2\text{O}(\text{l}) \quad \Delta h = -1411 \text{ kJ/mole}$
 $\text{C}_2\text{H}_6(\text{g}) + 7 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{l}) \quad \Delta h = -1560 \text{ kJ/mole}$
 $\text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l}) \quad \Delta h = -285 \text{ kJ/mole}$

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web thus for reaction 9 2 in which 1 mole of H_2O and 1 mole of OH^- combine to form 2 moles of H_2O $\Delta H = -55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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