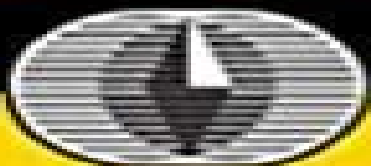


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calculated mass of the product expected $\times 100$ in this reaction a yield 80 is anticipated

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g $\text{Al} \times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 / 302.3 \text{ g Al}_2\text{O}_3 = 26.982 \text{ g Al}$ 4 mol $\text{Al} \times 1 \text{ mol Al}_2\text{O}_3$ c determine the percent yield

yield actual $\times 100$ $260.0 \text{ g} \times 100 / 86.01$ theoretical 302.3 g

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c is the limiting reagent 4 33 g of h 2 are left over 13

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yield $\frac{\text{theoretical yield}}{\text{actual yield}} \times 100$ based on this definition we would expect a percent yield

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