

# Practical Guide to Naming Oxyacids

- The names for polyatomic ions follow a similar set of rules to those for naming oxyacids, so the name of the polyatomic ion reveals the name of the acid.



Number of oxygens compared to most common	Polyatomic ion			Acid name	
	Example	Prefix	Suffix	Prefix	Suffix
Two Less	$\text{ClO}^-$	hypo	ite	hypo	ous
One Less	$\text{ClO}_2^-$	–	ite	–	ous
Same	$\text{ClO}_3^-$	–	ate	–	ic
One More	$\text{ClO}_4^-$	per	ate	per	ic

- Examples
  - $\text{ClO}^-$  is hypochlorite, so  $\text{HClO}$  is hypochlorous acid.
  - $\text{NO}_2^-$  is nitrite, so  $\text{HNO}_2$  is nitrous acid.

# Naming Acids And Bases Evan P

**Mike Jess**



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initial volume v1 10 liter 10 dm<sup>3</sup> 10 x 10<sup>-3</sup> m<sup>3</sup>

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