

AP Chemistry Review

Chapter 6 and 7, Atomic Structure and Periodicity

- Be familiar with the electromagnetic spectrum
 1. What types of light are shorter in length than microwave? Know the whole order of the spectrum.
 2. What type of light helps scientists to learn about electron structure?
 3. What creates the emission lines in a spectra for an element? What would cause only certain wavelengths of light to be absorbed by an atom?
- Convert between frequency and wavelength
 4. Calculate the frequency of light emitted by exciting Sr atoms if the wavelength is 461 nm.
 5. Calculate the wavelength of light emitted by exciting Na atoms if the frequency is $3.45 \times 10^{14} \text{ s}^{-1}$
- Calculate the energy of a wave or an electron
 6. Calculate the energy of 1 mole of photons emitted by a Li atom if the wavelength is 670nm.
- Define and understand:
 7. Ground state, excited state, Pauli exclusion principle, Heisenberg Uncertainty Principle, Hund's rule.
- Know the maximum number of electrons in and shape of each type of orbital.
 8. 3s 5p 4f 6d
- Following the Aufbau principle, determine electron configurations, both longhand and condensed, for elements and ions
- Match each explanation below with the correct letter
 - (A) $1s^2 2s^2 2p^5 3s^2 3p^5$
 - (B) $1s^2 2s^2 2p^6 3s^2 3p^6$
 - (C) $1s^2 2s^2 2p^6 2d^{10} 3s^2 3p^6$
 - (D) $1s^2 2s^2 2p^6 3s^2 3p^6 3d^5$
 - (E) $1s^2 2s^2 2p^6 3s^2 3p^6 3d^5 4s^2$
- 9. An impossible electron configuration
- 10. The ground-state configuration for the atoms of a transition element
- 11. The ground-state configuration of a negative ion of a halogen
- 12. The ground-state configuration of a common ion of an alkaline earth element
- 13. List 3 species that are isoelectronic
- 14. Which of the following represents the ground state electron configuration for the Mn^{3+} ion? (Atomic number Mn = 25)
 - (a) $1s^2 2s^2 2p^6 3s^2 3p^6 3d^4$
 - (b) $1s^2 2s^2 2p^6 3s^2 3p^6 3d^5 4s^2$
 - (c) $1s^2 2s^2 2p^6 3s^2 3p^6 3d^2 4s^2$
 - (d) $1s^2 2s^2 2p^6 3s^2 3p^6 3d^6 4s^2$
 - (e) $1s^2 2s^2 2p^6 3s^2 3p^6 3d^5 4s^1$
- Draw Dot diagrams for elements
 15. Boron Calcium Krypton

Merrill Chemistry Study Guide Chapter 6 Answers

JS Bruner



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Table of Contents Merrill Chemistry Study Guide Chapter 6 Answers

1. Understanding the eBook Merrill Chemistry Study Guide Chapter 6 Answers
 - The Rise of Digital Reading Merrill Chemistry Study Guide Chapter 6 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Merrill Chemistry Study Guide Chapter 6 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Merrill Chemistry Study Guide Chapter 6 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Merrill Chemistry Study Guide Chapter 6 Answers
 - Personalized Recommendations
 - Merrill Chemistry Study Guide Chapter 6 Answers User Reviews and Ratings
 - Merrill Chemistry Study Guide Chapter 6 Answers and Bestseller Lists
5. Accessing Merrill Chemistry Study Guide Chapter 6 Answers Free and Paid eBooks
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 - Merrill Chemistry Study Guide Chapter 6 Answers eBook Subscription Services
 - Merrill Chemistry Study Guide Chapter 6 Answers Budget-Friendly Options

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 - Merrill Chemistry Study Guide Chapter 6 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
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 - Highlighting and Note-Taking Merrill Chemistry Study Guide Chapter 6 Answers
 - Interactive Elements Merrill Chemistry Study Guide Chapter 6 Answers
8. Staying Engaged with Merrill Chemistry Study Guide Chapter 6 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Merrill Chemistry Study Guide Chapter 6 Answers
9. Balancing eBooks and Physical Books Merrill Chemistry Study Guide Chapter 6 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Merrill Chemistry Study Guide Chapter 6 Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Merrill Chemistry Study Guide Chapter 6 Answers
 - Setting Reading Goals Merrill Chemistry Study Guide Chapter 6 Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Merrill Chemistry Study Guide Chapter 6 Answers
 - Fact-Checking eBook Content of Merrill Chemistry Study Guide Chapter 6 Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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