

Name _____ Date _____ Class _____

MATH SKILLS TRANSPARENCY WORKSHEET

1

Interpreting and Drawing Graphs

Use with Chapter 2,
Section 2.4

1. What kind of graph is shown on the transparency?

The graph shown is a circle graph.

2. What does the circle represent?

The circle represents the total number of people aged 18–21 who responded to the survey.

3. Based on the data shown, what percentage of 18- to 21-year-olds think of themselves as Republicans? As Democrats?

19% think of themselves as Republicans; 25% think of themselves as Democrats.

4. Which group do most 18- to 21-year-olds say they belong to?

Most 18- to 21-year-olds (38%) say they are independent.

5. The data on the right shows the political affiliations for the general population. Make a graph that compares the political affiliations of 18- to 21-year-olds with those of the general population. Label the appropriate parts of your graph.

Students might make a bar graph that uses bars of two colors, one to represent 18- to 21-year-olds and one to represent the general population.

Political Affiliation	
Political affiliation	Percent of general population who identify with the affiliation
Republican	25
Democrat	39
Independent	36
Other party	0

Source: The Economist, July 14, 2000

Use the graph you made in question 5 to answer the following questions.

6. What kind of graph did you make?

Answers may vary, but a bar graph would be a clear way of presenting the comparison of data.

7. Compare the responses of the general population with those of 18- to 21-year-olds. How are they alike? How are they different?

Both groups are divided among affiliations in a similar manner: most say they are independent. However, a smaller portion of the general population says they are independent.

8. What is the greatest difference between the responses of the general population and those of 18- to 21-year-olds?

The greatest difference is the decrease in the percentage of people in the general population who say they are independent.

Name _____ Date _____ Class _____

MATH SKILLS TRANSPARENCY WORKSHEET

2

Visualizing the Conservation of Mass

Use with Chapter 2,
Section 2.2

1. How many potassium atoms are in the reactants of the chemical reaction shown? How

many are in the products? $2; 2$

2. How many oxygen atoms are in the reactants of the chemical reaction shown? How many

are in the products? $2; 2$

3. How many hydrogen atoms are in the reactants of the chemical reaction shown? How

many are in the products? $4; 4$

4. Assume that the chemical reaction shown started out having a total of 17 g of potassium and water. How much potassium hydroxide and hydrogen gas will be produced by the chemical reaction? Show your work.

$$\text{Mass}_{\text{reactants}} = \text{Mass}_{\text{products}}$$

If the reactants total 17 g, then the products will total 17g.

5. Assume that the chemical reaction shown started out having 4 atoms of potassium and 4 molecules of water. How many molecules of potassium hydroxide will be produced by the chemical reaction? How many hydrogen atoms will result?

$$\text{Mass}_{\text{reactants}} = \text{Mass}_{\text{products}}$$

There will be 4 molecules of potassium hydroxide and 12 atoms of hydrogen.

6. Assume that you are working with a chemical reaction that decomposes water into hydrogen and oxygen. You begin with 16 grams of water and end with 12 grams of oxygen. If all of the water decomposes, how many grams of hydrogen gas will result?

$$\text{Mass}_{\text{reactants}} = \text{Mass}_{\text{products}}$$

$$\text{Mass}_{\text{water}} = \text{Mass}_{\text{hydrogen}} + \text{Mass}_{\text{oxygen}}$$

$$\text{Mass}_{\text{hydrogen}} = \text{Mass}_{\text{water}} - \text{Mass}_{\text{oxygen}}$$

$$\text{Mass}_{\text{hydrogen}} = 16 \text{ g} - 12 \text{ g} = 4 \text{ g}$$

7. Assume that you are working with a chemical reaction that synthesizes salt from sodium and chlorine. You begin with 70.9 g of chlorine. You synthesize 114.90 grams of salt. If all of the reactants were used up, how many grams of sodium did you begin with?

$$\text{Mass}_{\text{reactants}} = \text{Mass}_{\text{products}}$$

$$\text{Mass}_{\text{sodium}} + \text{Mass}_{\text{chlorine}} = \text{Mass}_{\text{sodium chloride}}$$

$$\text{Mass}_{\text{sodium}} = \text{Mass}_{\text{sodium chloride}} - \text{Mass}_{\text{chlorine}}$$

$$\text{Mass}_{\text{sodium}} = 114.9 \text{ g} - 70.9 \text{ g} = 44 \text{ g}$$

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Larry Holt



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