

- a.) 101 Md Element named in honor of the "Father of the Periodic Table"
- b.) number The chemical and physical properties are a function of their atomic (mass / number).
- c.) period Term for horizontal rows
- d.) family group Term for vertical columns
- e.) P Element of 3rd period, $\frac{1}{2}$ filled p sublevel.
- f.) Mg Which element(s) has no d electrons? Mg, Au, I, Cs
- g.) 1, 0, Mg Number of elements in 3rd period with only 2 valence electrons
- h.) Halogens Family group that ends with $s^2 p^5$
- i.) Noble Gas Family with filled valence shell
- j.) He Element in family with filled valence shell that only contains s electrons.
- k.) ~~Se~~ Sb $[\text{Kr}]5s^2 4d^{10} 5p^3$
- l.) Ca 4th period alkaline Earth metal
- m.) 1 Number of valence electrons that the alkali metals have
- n.) Li lightest alkali metal
- o.) Fr largest alkali metal
- p.) Na largest element in 3rd period.
- q.) $4f^2, 3p^2, 2d^1$ Which electron configuration(s) is/are not allowed: $7p^2, 4f^6, 3p^2, 2d^1$
- r.) $3, 0, 1, -\frac{1}{2}$ 2 electrons in an orbital. If quantum set = $3, 0, 1, +\frac{1}{2}$ what is other quantum set of numbers?
- s.) Tc Ground state, 5th period element with valence orbital notation: $\uparrow\downarrow \quad \uparrow\downarrow \quad \uparrow \quad \uparrow$
- t.) 5f Sublevel filling in the actinide series
- u.) Bi element in 6th period with 5 valence electrons
- v.) 10 maximum number of electrons allowed in the 3d sublevel.
- w.) 12 How many s sublevel electrons does an atom of lead have? ($1s^2, 2s^2, 3s^2, 4s^2, 5s^2, 6s^2$)
- x.) 2 6 4 Number of valence electrons for Sr, S, Pb
- y.) 16 Number of orbitals for the 4th energy level $2n^2 = 2(4^2) = 32 e^- / 2 = 16 \text{ orbitals}$
- z.) Pauli Exclusion P. No 2 electrons in the same atom can have the same 4 quantum numbers.

2. Circle the larger particle

- a.) Mg or Ca
- b.) P or Se
- c.) K or As
- d.) Sn or S
- e.) Al^{3+} or Al

- f.) Ni or Ag
- g.) 1 or P
- h.) O^{2-} or S^{2-}
- i.) Ca^{2+} or Cl
- j.) As or Br

3. Circle particle with larger ionization energy

- a.) Se or S
- b.) K or Cs
- c.) F or O

- d.) O or P
- e.) Al or Ca
- f.) I or O

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1. Understanding the eBook Gibbs Energy Chemistry If8766 Answers
 - The Rise of Digital Reading Gibbs Energy Chemistry If8766 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Gibbs Energy Chemistry If8766 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gibbs Energy Chemistry If8766 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gibbs Energy Chemistry If8766 Answers
 - Personalized Recommendations
 - Gibbs Energy Chemistry If8766 Answers User Reviews and Ratings
 - Gibbs Energy Chemistry If8766 Answers and Bestseller Lists

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 - Gibbs Energy Chemistry If8766 Answers Public Domain eBooks
 - Gibbs Energy Chemistry If8766 Answers eBook Subscription Services
 - Gibbs Energy Chemistry If8766 Answers Budget-Friendly Options
6. Navigating Gibbs Energy Chemistry If8766 Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Gibbs Energy Chemistry If8766 Answers Compatibility with Devices
 - Gibbs Energy Chemistry If8766 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gibbs Energy Chemistry If8766 Answers
 - Highlighting and Note-Taking Gibbs Energy Chemistry If8766 Answers
 - Interactive Elements Gibbs Energy Chemistry If8766 Answers
8. Staying Engaged with Gibbs Energy Chemistry If8766 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gibbs Energy Chemistry If8766 Answers
9. Balancing eBooks and Physical Books Gibbs Energy Chemistry If8766 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gibbs Energy Chemistry If8766 Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Gibbs Energy Chemistry If8766 Answers
 - Setting Reading Goals Gibbs Energy Chemistry If8766 Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gibbs Energy Chemistry If8766 Answers
 - Fact-Checking eBook Content of Gibbs Energy Chemistry If8766 Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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