

Characteristics

Answer all (except (a)(ii)) Questions

- (a) (i) When (aq) is used as a primary standard, what must its concentration be?
- (ii) When is the standard solution used in the titration?
- (iii) When is the pH of the solution measured?
- (iv) When is a standard solution used and how does it compare to a standard solution? When is the pH of the solution measured?
- (v) At the equivalence point of a titration, the pH is _____.
- (vi) At the equivalence point of a titration, the pH is _____.
- (vii) When is the pH measured?
- (viii) When does the pH of the solution change?
- (ix) When does the pH of the solution change? When is the pH of the solution measured?

Use the data in the table to answer the following questions

Compound	Concentration (mol/L)	pH	Volume (mL)	Volume (mL)
0.100 M HCl	0.100	1.0	10.0	10.0
0.100 M NaOH	0.100	13.0	10.0	10.0
0.100 M H ₂ SO ₄	0.100	1.0	10.0	10.0
0.100 M Na ₂ SO ₄	0.100	13.0	10.0	10.0
0.100 M H ₃ PO ₄	0.100	1.0	10.0	10.0
0.100 M Na ₃ PO ₄	0.100	13.0	10.0	10.0

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