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Embarq User Guide

Wolfgang Guggemos

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web Ksp of calcium hydroxide 1 author J M McCormick last update October 13 2013 introduction an equilibrium constant K is related to ΔG for a process through eqn 1 and as long as ΔH and ΔS are independent of temperature eqn 1 can be combined with eqn 2 to determine ΔH and ΔS from the temperature dependence of K

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web in today's experiment you will determine the solubility product ksp of calcium hydroxide Ca(OH)_2 by measuring the concentration of Ca(OH)_2 in a saturated solution calcium hydroxide is a sparingly soluble salt that dissolves according to the following reaction the solubility product expression for this reaction is

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web expert verified chem 182 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water a saturated aqueous solution of Ca(OH)_2 is represented in equation form as shown below $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$ the solubility product expression describes in mathematical terms the equilibrium

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web the chemical equation used was $\text{Ca(OH)}_2 \text{ s} + \text{Ca}^{2+} \text{ aq} + 2\text{OH}^- \text{ aq}$ the solubility product was determined to be $K_{\text{sp}} \text{ Ca}^{2+} \text{ OH}^-$ from the titration which was made by dripping a concentrated solution of calcium hydroxide into hydrochloric acid the calcium hydroxide was filtered before starting the titration

18 1 solubility product constant K_{sp} chemistry libretexts - Feb 01 2023

web jul 12 2023 because the concentration of a pure solid such as $\text{Ca}_3(\text{PO}_4)_2$ is a constant it does not appear explicitly in the equilibrium constant expression the equilibrium constant expression for the dissolution of calcium phosphate is therefore $K_{\text{sp}} \text{ Ca}^{2+}_3 \text{ PO}_4^{3-}_2$

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dissolves to

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we will filter the solid calcium hydroxide from the solution and then titrate the dissolved calcium hydroxide in solution with a known concentration of hydrochloric acid using titrimetric calculations we will determine the Ca^{2+} and OH^- concentrations in solution

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we calculate the ksp for calcium hydroxide $\text{ksp Ca(OH)}_2 = 5 \times 10^{-6}$ find the accepted value of the ksp for calcium hydroxide and compare it with your value discuss the discrepancy and suggest possible sources of experimental error
accepted value for ksp for sodium hydroxide is 5×10^{-6}