

Question	Answer	Marks	Part marks
1(a)(i)	$600 \div (11 + 9) \times 11$ [=330] with no errors seen	M1	Could be in separate steps
1(a)(ii)	270	1	
1(b)(i)	372 <i>cao nfw/w</i>	3	B2 for answer 371.7... or M1 for $330 \times \left(1 + \frac{1.5}{100}\right)^8$ <i>oe</i> not spoiled After zero scored, SC1 for answer 42 or 41.7...
1(b)(ii)	12.6 or 12.7 or 12.63 to 12.73	2	M1 for $\frac{\text{their (b)(i)} - 330}{330}$ or $\frac{\text{their (b)(i)}}{330} \times 100$ <i>soi</i> by 112.7 or 113 After zero scored, SC1 for answer 12%
1(c)(i)	$\frac{99}{280}$ <i>cao final answer</i>	1	
1(c)(ii)	27.5[0]	3	M2 for $24.75 + \frac{100 - 10}{100}$ <i>oe</i> or M1 for recognising 24.75 as 90[%] <i>oe</i>
1(d)(i)	32 <i>cao</i>	2	M1 for $\left(1 - \frac{20}{100}\right)\left(1 - \frac{15}{100}\right)[x]$ <i>oe</i> or for $0.15 \times 0.8 [x]$ <i>oe</i>
1(d)(ii)	13 <i>cao</i>	2	M1 for $\left(1 - \frac{20}{100}\right)\left(1 - \frac{15}{100}\right) \times x = 40.84 - 32$ <i>oe seen</i> or for $\text{their (d)(i)} + \left(1 - \left(\frac{\text{their (d)(i)}}{100}\right)\right)x = 40.84$ <i>oe</i>
2(a)(i)	Image at (8, 1), (10, 5), (8, 5)	2	B1 for translation $\begin{pmatrix} 6 \\ k \end{pmatrix}$ or $\begin{pmatrix} k \\ -5 \end{pmatrix}$ or 3 correct points not joined
2(a)(ii)	Image at (4, 10), (4, 8), (8, 8)	2	B1 for rotation 90° anticlockwise but different centre or for rotation 90° clockwise about (-4, 10) or 3 correct points not joined
2(a)(iii)	Image at (6, 3), (6, 5), (7, 5)	2	B1 for enlargement factor $\frac{1}{2}$ but incorrect centre or 3 correct points not joined
2(b)	Reflection	1	
	$y = -x$ <i>oe</i>	1	If zero scored, M1 for correct use of matrix product

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web information solubility rules a solute is considered soluble if an appreciable amount of it can be dissolved in a given amount of the solvent for example both table salt NaCl and table sugar $\text{C}_{12}\text{H}_{22}\text{O}_{11}$ are soluble substances in water a solute is considered insoluble if very little of it dissolves in a given amount of the solvent

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web predict what might happen to the solubility of the insoluble salt if a 1.0 M acid solution was added drop wise to the beaker hint which ion species in the beaker will react with the acid how will this reaction alter the concentration of the common ion