

NAMING BINARY COMPOUNDS (IONIC)

Name _____

Name the following ionic compounds using Roman numerals where necessary.

1. BaCl_2 _____

2. NaF _____

3. Ag_2O _____

4. CuBr _____

5. CuBr_2 _____

6. FeO _____

7. Fe_2O_3 _____

8. MgS _____

9. Al_2O_3 _____

10. CaI_2 _____

11. K_2S _____

12. CrCl_2 _____

13. CrCl_3 _____

14. CaO _____

15. Ba_3P_2 _____

16. Hg_2I_2 _____

17. Na_2O _____

18. BeS _____

19. MnO _____

20. Mn_2O_3 _____

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Glencoe Science McGraw-Hill Staff, 2001-06 **Prentice Hall Chemistry**, 2000 **Chemistry** John S. Phillips, Cheryl Wistrom, 2000 **Science Software Quarterly**, 1985 **Master Analytical Manual: Ionic methods** Oak Ridge National Laboratory. Analytical Chemistry Division, 1959 **Chemistry, Student Study Guide** James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement **The Success Manual for General Chemistry** Elizabeth Kean, Catherine Middlecamp, 1986 *Chemical Principles* Peter Atkins, Loretta

Jones,2007-08 Written for calculus inclusive general chemistry courses Chemical Principles helps students develop chemical insight by showing the connections between fundamental chemical ideas and their applications Unlike other texts it begins with a detailed picture of the atom then builds toward chemistry's frontier continually demonstrating how to solve problems think about nature and matter and visualize chemical concepts as working chemists do Flexibility in level is crucial and is largely established through clearly labeling separating in boxes the calculus coverage in the text Instructors have the option of whether to incorporate calculus in the coverage of topics The multimedia integration of Chemical Principles is more deeply established than any other text for this course Through the unique eBook the comprehensive Chemistry Portal Living Graph icons that connect the text to the Web and a complete set of animations students can take full advantage of the wealth of resources available to them to help them learn and gain a deeper understanding **How to Survive (and Even Excel In)**

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