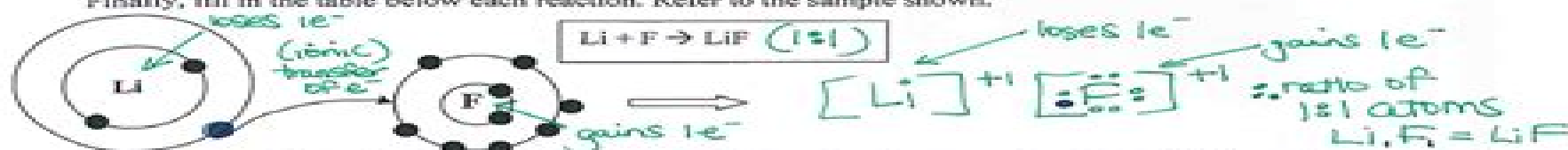


Name Solutions

Date _____ Period _____

Ionic Bonding Worksheet

For each pair of elements below draw an atomic diagram showing electrons in different energy levels. Draw arrows to show where the outer electrons will go during a chemical reaction, then draw the resulting compound. Finally, fill in the table below each reaction. Refer to the sample shown.



Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
Li	1 ve	transfers 1ve ⁻ to F.	Li ⁺
F	7 ve	accepts 1ve ⁻ from Li	F ⁻

Reactions	Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
1) $\text{Li} + \text{Cl} \Rightarrow \text{LiCl}$ 	Li	1 ve	donate 1e ⁻	Li ⁺
	Cl	1 ve	accepts 1e ⁻	Cl ⁻
2) $\text{Ca} + \text{O} \Rightarrow \text{CaO}$ 	Ca	2 ve	donates 2e ⁻	Ca ⁺²
	O	6 ve	accepts 2e ⁻	O ⁻²
3) $\text{Be} + \text{F} \Rightarrow \text{BeF}_2$ 	Be	2 ve	donates 2e ⁻ to each F.	Be ⁺²
	F	7 ve	each F accepts 1ve each	F ⁻¹
4) $\text{Mg} + \text{S} \Rightarrow \text{MgS}$ 	Mg	2 ve	donates 2e ⁻ to S	Mg ⁺²
	S	6 ve	accepts 2e ⁻	S ⁻²
5) $\text{K} + \text{F} \Rightarrow \text{KF}$ 	K	1 ve	donates 1ve ⁻	K ⁺¹
	F	7 ve	accepts 1ve ⁻	F ⁻¹

Ionic Bonding Chemquest 20 Answer

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