

Naming Ionic Compounds

What are the structural units that make up ionic compounds and how are they named?

Why?

When working in chemistry, it is often convenient to write a chemical in symbols. For example we might write down the substance table salt as NaCl. In talking about chemistry however, it is a bit tacky to say "en-ay see-ell" when we want to refer to a substance. Also, in formal writing we should use the name of the compound rather than its symbols. Therefore we need to learn how to say the proper names of ionic substances.

Model 1 – Ion Charges for Selected Elements

1	H ⁺															
2	Li ⁺	Be ²⁺	Transition elements						Al ³⁺		N ³⁻	O ²⁻	F ⁻			
3	Na ⁺	Mg ²⁺									P ³⁻	S ²⁻	Cl ⁻			
4	K ⁺	Ca ²⁺	Fe ²⁺ Fe ³⁺	Ni ²⁺ Ni ³⁺	Cu ⁺ Cu ²⁺	Zn ²⁺								Br ⁻		
5	Rb ⁺	Sr ²⁺			Ag ¹⁺						Sn ²⁺ Sn ⁴⁺			I ⁻		
6		Ba ²⁺				Hg ₂ ²⁺ Hg ²⁺				Pb ²⁺ Pb ⁴⁺						
← Cations →																

1. Based on the information in Model 1:

- a. Identify three elements that form only one cation. $\leftarrow$ Positive Ion
 Na^+ , Sr^{2+} , Al^{3+} (Any box with one element works.)
- b. Identify three elements that form only one anion. $\leftarrow$ Negative Ion
 P^{3-} , O^{2-} , Br^- (Any box with one element works.)
- c. Identify three elements that form more than one cation.
 Fe^{+2} and Fe^{+3} , Ni^{+2} and Ni^{+3} , Cu^{+1} and Cu^{+2}
- d. In what region of the periodic table are these "multiple ion" elements usually located?
Transition elements or "d" block

2. Consider the ions of potassium (K) and sulfur (S). Write chemical formulas for all possible ionic compounds involving these ions, using the simplest ratio(s) of potassium (K) and sulfur (S). Keep in mind that the sum of the charges in an ionic compound must equal zero.
- $2K^+$ and S^{2-} or K_2S
3. Consider the ions of iron (Fe) and sulfur (S). Write chemical formulas for all possible ionic compounds involving these ions, using the simplest ratio(s) of iron (Fe) and sulfur (S). Keep in mind that the sum of the charges in an ionic compound must equal zero.



Ionic Compounds Study Guide Answer Sheet

John Giorgi



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Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter **Wastewater Operator**

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