



# CHEMICAL BONDS



## IONIC, COVALENT, & METALLIC

**FORMS OF CHEMICAL BONDS**  
Bonding: combination of atoms

Atoms combine to form molecules or compounds by sharing or transferring electrons. This process is called bonding. The resulting molecule or compound has properties different from the individual atoms.

**Types of Bonds:**

- Covalent Bond:** Atoms share electrons to form a bond. Example:  $H_2$ ,  $H_2O$ ,  $CO_2$ .
- Ionic Bond:** One atom transfers electrons to another atom. Example:  $NaCl$ ,  $MgO$ .
- Metallic Bond:** Atoms in a metal share their valence electrons in a "sea of electrons". Example:  $Fe$ ,  $Cu$ ,  $Al$ .

**Covalent, Ionic, and Metallic Bonds**

Covalent bonds are formed by sharing electrons between two non-metal atoms. Ionic bonds are formed by the transfer of electrons from a metal atom to a non-metal atom. Metallic bonds are formed by the sharing of valence electrons among many metal atoms.

**Properties of Bonds:**

- Covalent bonds are strong and directional.
- Ionic bonds are strong and non-directional.
- Metallic bonds are strong and non-directional.

**Diagram:** A periodic table of elements showing the general location of elements that form covalent, ionic, and metallic bonds.



**IONIC BONDING**

Atoms combine to form ionic compounds by transferring electrons. The resulting compound has properties different from the individual atoms.

**Diagram:** A diagram showing the transfer of electrons from a metal atom to a non-metal atom, resulting in an ionic bond.

**Properties of Ionic Compounds:**

- High melting and boiling points.
- Brittle.
- Good conductors of electricity when dissolved in water.

**Diagram:** A diagram showing the structure of an ionic compound, with positive and negative ions arranged in a crystal lattice.

**COVALENT BONDING**

Atoms combine to form covalent compounds by sharing electrons. The resulting compound has properties different from the individual atoms.

**Diagram:** A diagram showing the sharing of electrons between two non-metal atoms, resulting in a covalent bond.

**Properties of Covalent Compounds:**

- Low melting and boiling points.
- Soft.
- Poor conductors of electricity.

**Diagram:** A diagram showing the structure of a covalent compound, with atoms sharing electrons in a network.

**METALLIC BONDING**

Atoms in a metal share their valence electrons in a "sea of electrons". The resulting metal has properties different from the individual atoms.

**Diagram:** A diagram showing the structure of a metal, with atoms and a sea of electrons.

**Properties of Metals:**

- High melting and boiling points.
- Malleable and ductile.
- Good conductors of electricity.

## GUIDED READING ACTIVITY

# Ionic And Metallic Bonding Guided Reading

**Deepa Chaudhary,Virendra,Priyanka  
Sharm,Rashmi Sharma,Ashwini Kumar  
Prasad**

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