

HUMAN
RIGHTS
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WORLD REPORT

2008

EVENTS OF 2007

Human Right Watch Report 2008

David Kirk



Human Right Watch Report 2008:

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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu}(\text{NO}_3)_2 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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$$8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} = 8.3 \text{ mole H}_2\text{O}$$

 8.3 mole h2s x 18g h2o i mole h2o 149 4g h2o is your theoretical yield so percent yield 137 1g h2o actual yield 149 4g h2o x 100
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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$

 if the reaction of 21 3 grams of o2 produces a 47 3 yield how many grams of co would be produced step one figure out how many moles of the limiting reagent you have

$$\frac{21.3 \text{ g O}_2}{32 \text{ g/mol}} = 0.666 \text{ moles O}_2$$

 0 66 moles 32 g m

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calculated mass of the product expected $\times 100$ in this reaction a yield 80 is anticipated

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g Al $\times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 / 302.3 \text{ g Al}_2\text{O}_3 = 26.982 \text{ g Al}$ 4 mol Al 1 mol Al_2O_3 c determine the percent yield
yield actual $\times 100$ 260.0 g $\times 100$ 86.01 theoretical 302.3 g

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C is the limiting reagent 4.33 g of H₂ are left over 13

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 $80 \text{ g I}_2\text{O}_5 \times \frac{1 \text{ mol I}_2\text{O}_5}{1 \text{ mol I}_2} \times \frac{1 \text{ mol I}_2}{1 \text{ mol I}_2\text{O}_5} \times \frac{333.8 \text{ g I}_2}{1 \text{ mol I}_2} = 333.8 \text{ g I}_2$
 $28 \text{ g CO} \times \frac{1 \text{ mol CO}}{1 \text{ mol CO}} = 28 \text{ g CO}$

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