

Experiment 1

1. Write the general equation for an equilibrium/phase transition and give the conditions of each. Example: melting

Experiment 2: Liquid and Vapor Equilibrium in a Closed System

In a closed system, equilibrium is established between the liquid and the vapor of a substance. The conditions of equilibrium are:

- a. pressure is equal to the external pressure (atmospheric pressure)

a. $p_{\text{liquid}} = p_{\text{vapor}} = p_{\text{external}}$	b. $p_{\text{liquid}} = p_{\text{vapor}} = p_{\text{external}}$
c. $p_{\text{liquid}} = p_{\text{vapor}} = p_{\text{external}}$	d. $p_{\text{liquid}} = p_{\text{vapor}} = p_{\text{external}}$

Experiment 3: Vapor Pressure of a Pure Substance

- a. The vapor pressure of a pure substance is the pressure exerted by the vapor of the substance in equilibrium with the liquid phase.

b. What is the relationship between the vapor pressure and the temperature?

c. What is the relationship between the vapor pressure and the boiling point?

d. What is the relationship between the vapor pressure and the boiling point?

e. What is the relationship between the vapor pressure and the boiling point?

- f. In a closed system, equilibrium is established between the liquid and the vapor of a pure substance. The conditions of equilibrium are:

a. pressure

b. pressure

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Tobias Bleicker



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