

# Honors Chemistry Worksheet 3

## Stoichiometry Practice Problems

Name \_\_\_\_\_ Period \_\_\_\_\_ Date \_\_\_\_\_

Instructions: Balance the following chemical equations and then determine the missing information for each of the conditions given. The four questions related to each equation are independent of one another. Answers for a particular numbered problem should be in the same units as the information given (i.e. grams to grams, moles to moles, particles to particles).



|    |             |  |       |  |       |
|----|-------------|--|-------|--|-------|
| 1. | <u>5mol</u> |  | _____ |  | _____ |
|----|-------------|--|-------|--|-------|

|    |       |  |               |  |       |
|----|-------|--|---------------|--|-------|
| 2. | _____ |  | <u>2.5mol</u> |  | _____ |
|----|-------|--|---------------|--|-------|

|    |            |  |       |  |       |
|----|------------|--|-------|--|-------|
| 3. | <u>25g</u> |  | _____ |  | _____ |
|----|------------|--|-------|--|-------|

|    |       |  |       |  |                                                   |
|----|-------|--|-------|--|---------------------------------------------------|
| 4. | _____ |  | _____ |  | <u><math>2.41 \times 10^{24}</math> molecules</u> |
|----|-------|--|-------|--|---------------------------------------------------|



|    |       |  |               |  |       |  |       |
|----|-------|--|---------------|--|-------|--|-------|
| 1. | _____ |  | <u>.75mol</u> |  | _____ |  | _____ |
|----|-------|--|---------------|--|-------|--|-------|

|    |            |  |       |  |       |  |       |
|----|------------|--|-------|--|-------|--|-------|
| 2. | <u>50g</u> |  | _____ |  | _____ |  | _____ |
|----|------------|--|-------|--|-------|--|-------|

|    |       |  |       |  |               |  |       |
|----|-------|--|-------|--|---------------|--|-------|
| 3. | _____ |  | _____ |  | <u>3.5mol</u> |  | _____ |
|----|-------|--|-------|--|---------------|--|-------|

|    |       |  |       |  |       |  |                                                    |
|----|-------|--|-------|--|-------|--|----------------------------------------------------|
| 4. | _____ |  | _____ |  | _____ |  | <u><math>1.204 \times 10^{23}</math> molecules</u> |
|----|-------|--|-------|--|-------|--|----------------------------------------------------|



|    |                                               |  |       |  |       |
|----|-----------------------------------------------|--|-------|--|-------|
| 1. | <u><math>1.51 \times 10^{23}</math> atoms</u> |  | _____ |  | _____ |
|----|-----------------------------------------------|--|-------|--|-------|

|    |       |  |                |  |       |
|----|-------|--|----------------|--|-------|
| 2. | _____ |  | <u>1.25mol</u> |  | _____ |
|----|-------|--|----------------|--|-------|

|    |       |  |       |  |                 |
|----|-------|--|-------|--|-----------------|
| 3. | _____ |  | _____ |  | <u>25 grams</u> |
|----|-------|--|-------|--|-----------------|

# Honors Chemistry Stoichiometry Test Answers

**Chao Zhang**



## Honors Chemistry Stoichiometry Test Answers:

Tests Richard C. Sweetland, Daniel J. Keyser, 1983 Descriptions of over 3000 tests in English intended as a guide for psychologists educators and other personnel who need test information to meet their assessment needs Topical arrangement under 3 main sections of psychology education and business Each entry gives test name associated personal names intended age group purpose description time range scoring cost and publisher Indexes by test titles authors publishers visually impaired tests and scoring services The Science Teacher, 1982 Some issues are accompanied by a CD ROM on a selected topic **The ETS Test Collection Catalog** Educational Testing Service. Test Collection, 1993 The major source of information on the availability of standardized tests Wilson Library Bulletin Covers commercially available standardized tests and hard to locate research instruments **The Pharmaceutical Era**, 1905 **A Whole Year of Chemistry Quizzes** Siegfried Howell, 2019-12-12 This book A Whole Year of Chemistry Quizzes was written to provide easy to use and grade quizzes to assess the comprehension of honors students Advance Placement students AP and International Baccalaureate IB students This book of quizzes guides the teacher and the student through what is required in a non watered down chemistry course that leads students towards test and college readiness The outline of this book has a minimum of 4 quizzes per chapter that prepares students for the formative assessment associated at the end of all chapters The 25 chapters include topics that are covered in the honors chemistry setting as well as specialty topics like thermodynamics kinetics rates of reactions that are seen in the Advance Placement classes Included within this book are quizzes for the International Baccalaureate teacher that wishes to test students on environmental chemistry as well as biological and food chemistry This is a book that was written to fill the void of valuable resources needed for novice and experienced teachers in institutions that continually push for more summative assessments higher DOKs and rapid feedback while limiting preparation time As a teacher for over 25 years I know that any well outlined structured and comprehensive resource saves time in additional planning searching and preparing Use this book to help you identify and test students on topics that are important to their comprehension and success with their final test Chapter 1 Matter and change Chapter 2 measurement and calculations Chapter 3 Atoms The building blocks of matter Chapter 4 Arrangement of electrons in atoms Chapter 5 The periodic law Chapter 6 Chemical bonding Chapter 7 Chemical formulas and chemical compounds Chapter 8 Chemical equations and reactions Chapter 9 Stoichiometry Chapter 10 Physical characteristics of gases Chapter 11 Molecular composition of gases Chapter 12 Liquids and solids Chapter 13 Solutions Chapter 14 Ions in aqueous solution and colligative properties Chapter 15 Acids and bases Chapter 16 Acid base titrations Chapter 17 Reaction energy and reaction kinetics Chapter 18 Chemical equilibrium Chapter 19 Oxidation reduction reactions Chapter 20 Chemical thermodynamics Chapter 21 Carbon and hydrocarbons Chapter 22 Other organic compounds Chapter 23 Nuclear chemistry Chapter 24 Biological and Food chemistry Chapter 25 Environmental chemistry A Concrete Stoichiometry Unit

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