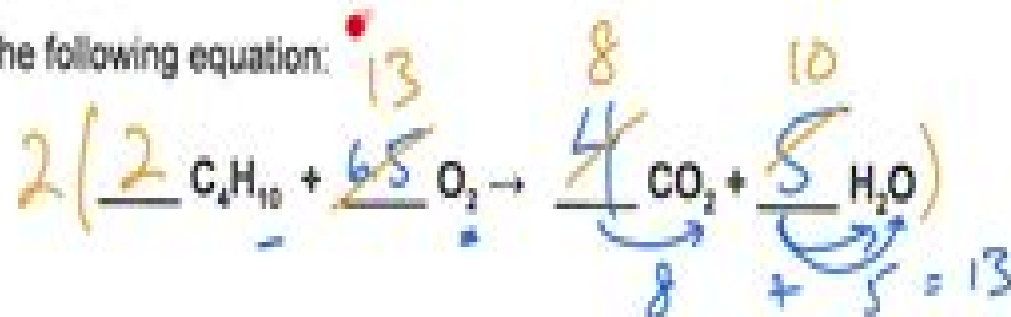


Balance the following equation:



1. How many moles are in 4.52×10^{30} molecules of O_2 ?

$$4.52 \times 10^{30} \text{ molecules} \times \frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ molecules}} = 7.51 \times 10^6 \text{ mol O}_2$$

2. If you have 3.5 moles of O_2 , how many moles of H_2O will you make?

$$3.5 \text{ mol O}_2 \times \frac{10 \text{ mol H}_2\text{O}}{13 \text{ mol O}_2} = \text{mol H}_2\text{O}$$

3. If you have 50.0 g of C_4H_{10} , how many moles of O_2 will you need?

4. If you have 100.0 g of O_2 , how many grams of CO_2 will you make?

Honors Chemistry Practice Test Semester 2

Christian G. Meyer



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