16 • Chemical Equilibrium

Consider the reaction system,

 $CoO(s) + H_2(g) \stackrel{\sim}{\leftarrow} Co(s) + H_2O(g)$. The equilibrium constant expression is

[CoO][H.] [Co][H.O]

- [H.O]
- $[Co][H_2O]$ [CoO][H.] Col[H.O] HI.
- Given the equilibrium.

2SO₂(g) + O₂(g) ∓ 2SO₂(g), if this equilibrium

is established by beginning with equal number of moles of SO₂ and O₂ in a 1.0 Liter bulb, then

the following must be true at equilibrium:

- a) [SO₂] = [SO₃]d) [SO₂] < [O₂]
- (b) $2|SO_2| = 2|SO_3|$ e) $|SO_2| > |O_2|$
 - c) [SO₂] = [O₂]

Questions 3 & 4 refer to the following:

At a given temperature, 0.300 mole NO, 0.200 mol Cb and 0.500 mol CINO were placed in a 25.0 Liter container. The following equilibrium is established: 2CINO(g) ₹ 2NO(g) + Cb(g)

- At equilibrium, 0.600 mol of CINO was present. The number of moles of Cb present at equilibrium is
 - cal 0.050 b) 0.100

e) 0.250

c) 0.150

- PRACTICE TEST
- The equilibrium constant, Kc, is:
 - a) 4.45 x 10⁻⁴
- d) 0.167
- b) 6.67 x 10⁻⁴
- e) 1500

- ch 0.111
- At 985°C, the equilibrium constant for the reaction.

$$H_2(g) + CO_2(g) \rightleftharpoons H_2O(g) + CO(g)$$

is 1.63. What is the equilibrium constant for the reverse reaction?

a) 1.63

d) 0.613

b) 0.815

e)) 1.00

- c) 2.66
- What is the relationship between K_p and K_c for the reaction, $2ICI(g) \rightleftharpoons I_2(g) + Cb(g)$?
 - (a) $k_p = K_c(RT)^{-1}$ (d) $K_p = K_c$

- b) $K_p = K_c(RT)$ e) $K_p = K_c(2RT)$
- c) $K_p = K_c(RT)^2$
- For the reaction 2NO₂(g) ₹ N₂O₄(g), K_n at 25°C is 7.3, when all partial pressures are expressed in atmospheres. What is Ke for this reaction? [R=0.0821 Latmmol* K*]

4270

d) 179

b) 0.0119

e) 2.06

c) = 0.291

Honors Chemistry Equilibrium Practice

Christian G. Meyer

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