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hot object is 60 c the specific heat of copper and water are 0.09 cal/gm c and 1 cal/gm c respectively

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web calorimetry problems calorimetry problems $q_{\text{sur}} = m \times c \times \Delta t$ q heat m mass $q_{\text{rxn}} = q_{\text{sur}}$ t_{f} t_{i} name per date c specific heat for water 4.184 J/g°C 1 what is the specific heat of aluminum if the temperature of a 28.4 g sample of aluminum is increased by 8.1 °C when 207 J of heat is added 2

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web use the information in this chart to answer the following questions 1 if i have 125 grams of water at a temperature of 25.0 °C and heat it to 145.0 °C what will the change in enthalpy be 2 sketch the heating curve temperature vs time for the process in 1.3 i have 18.7 grams of an unknown liquid chemical at a temperature of 35.0 °C

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web choose 1 answer the final temperature is closer to t_1 than to t_2 a the final temperature is closer to t_1 than to t_2 the final temperature is exactly halfway between t_1 and t_2 b the final temperature is exactly halfway between t_1 and t_2 the final temperature is closer to t_2 than to t_1 c

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web 6 a 218 g sample of steam at 121 $^{\circ}\text{C}$ is cooled to ice at 14 $^{\circ}\text{C}$ find the change in heat content of the system 7 in going from ice at 34 $^{\circ}\text{C}$ to steam at 138 $^{\circ}\text{C}$ a sample of water absorbs $1.41 \times 10^5 \text{ J}$ find the mass of the sample answers 1 2 68 $\times 10^4 \text{ J}$ 2 6 23 $\times 10^4 \text{ J}$ 3 1 71 $\times 10^4 \text{ J}$ 4 1 11 $\times 10^5 \text{ J}$ 5 2 11 $\times 10^5 \text{ J}$ 6

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