

Name: Townsend Smith

Date: _____

Student Exploration: Stoichiometry

Vocabulary: Avogadro's number, balanced equation, cancel, coefficient, conversion factor, dimensional analysis, formula mass, molar mass, mole, molecular mass, stoichiometry

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

1. A 250 mL glass of orange juice contains 22 grams of sugar. How much sugar is in a two-liter (2,000 mL) bottle of orange juice? 176g of sugar
2. It requires two sticks of butter to make a batch of 20 cookies. How much butter will it take to make 150 cookies? 15 butter sticks

Gizmo Warm-up

Just as a cook follows a recipe to decide how much of each ingredient to add, a chemist uses **stoichiometry** to determine the amounts of substances involved in chemical reactions. The *Stoichiometry Gizmo* allows you to try your hand at figuring out the amounts of reactants and products that take part in a chemical reaction.

To begin, check that this equation is shown:



1. Look at the **coefficients** (such as the "3" in 3CO) in front of each substance in the equation. The coefficients tell you how many molecules or atoms take part in a chemical reaction. In the spaces below, list the number of each molecule or atom in the equation:

Fe₂O₃ 1CO 3Fe 2CO₂ 3

2. In a **balanced equation**, the same number of each kind of atom is shown on each side of the equation. Calculate the number of iron (Fe), oxygen (O), and carbon atoms (C).

Reactants Iron: 2Oxygen: 6Carbon: 3Products Iron: 2Oxygen: 6Carbon: 3

Based on these values, is the equation balanced? Yes

New equation



New question

How many moles of carbon monoxide (CO) are required to react completely with 1.75 moles of iron (III) oxide (Fe₂O₃)?

Problems solved: 0



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1. Understanding the eBook Genetic Engineering Answer Key 13 Study Guide
 - The Rise of Digital Reading Genetic Engineering Answer Key 13 Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Genetic Engineering Answer Key 13 Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Genetic Engineering Answer Key 13 Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Genetic Engineering Answer Key 13 Study Guide
 - Personalized Recommendations
 - Genetic Engineering Answer Key 13 Study Guide User Reviews and Ratings
 - Genetic Engineering Answer Key 13 Study Guide and Bestseller Lists

5. Accessing Genetic Engineering Answer Key 13 Study Guide Free and Paid eBooks
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 - Genetic Engineering Answer Key 13 Study Guide eBook Subscription Services
 - Genetic Engineering Answer Key 13 Study Guide Budget-Friendly Options
6. Navigating Genetic Engineering Answer Key 13 Study Guide eBook Formats
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 - Adjustable Fonts and Text Sizes of Genetic Engineering Answer Key 13 Study Guide
 - Highlighting and Note-Taking Genetic Engineering Answer Key 13 Study Guide
 - Interactive Elements Genetic Engineering Answer Key 13 Study Guide
8. Staying Engaged with Genetic Engineering Answer Key 13 Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Genetic Engineering Answer Key 13 Study Guide
9. Balancing eBooks and Physical Books Genetic Engineering Answer Key 13 Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Genetic Engineering Answer Key 13 Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Genetic Engineering Answer Key 13 Study Guide
 - Setting Reading Goals Genetic Engineering Answer Key 13 Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Genetic Engineering Answer Key 13 Study Guide
 - Fact-Checking eBook Content of Genetic Engineering Answer Key 13 Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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