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web these ions are called spectator ions because they do not participate in the actual reaction canceling the spectator ions gives the net ionic equation which shows only those species that participate in the chemical reaction $2\text{Ag}^+ \text{aq}$ $\text{Cr}_2\text{O}_7^{2-} \text{aq}$ H_2O $\text{Cr}^{3+} \text{aq}$ $\text{Ag}_2\text{CrO}_4 \text{s}$

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web chemical equation in which all dissolved ionic reactants and products including spectator ions are explicitly represented by formulas for their dissociated ions molecular equation chemical equation in which all reactants and products are

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web jan 29 2020 the ionic form of this reaction is $2\text{Na}^+ \text{aq}$ $2\text{Cl}^- \text{aq}$ $\text{Cu}^{2+} \text{aq}$ $\text{SO}_4^{2-} \text{aq}$ $2\text{Na}^+ \text{aq}$ $\text{SO}_4^{2-} \text{aq}$ $\text{CuCl}_2 \text{s}$ the sodium ions and sulfate ion are the spectator ions in this reaction they appear unchanged in both the product and reactant side of the equation

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web mar 11 2014 solution ionic equation $\text{H}^+ \text{aq}$ $\text{Cl}^- \text{aq}$ $\text{Na}^+ \text{aq}$ $\text{HCO}_3^- \text{aq}$ $\text{Na}^+ \text{aq}$ $\text{Cl}^- \text{aq}$ H_2O $\text{CO}_2 \text{g}$ the spectator ions are Cl^- and Na^+ net ionic equation $\text{H}^+ \text{aq}$ $\text{Cl}^- \text{aq}$ $\text{Na}^+ \text{aq}$ $\text{HCO}_3^- \text{aq}$ $\text{Na}^+ \text{aq}$ $\text{Cl}^- \text{aq}$ H_2O $\text{CO}_2 \text{g}$ $\text{H}^+ \text{aq}$ $\text{HCO}_3^- \text{aq}$ H_2O $\text{CO}_2 \text{g}$ here is a video on spectator ions

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web the six Na^+ ions and the six Cl^- ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation $\text{Ba}^{2+} + 2\text{Cl}^- \rightarrow \text{BaCl}_2$

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web these spectator ions whose presence is required to maintain charge neutrality are neither chemically nor physically changed by the process and so they may be eliminated from the equation to yield a more succinct representation called a

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web video answer you know spectator ions are the ions present on the both sides of the equation indicate cancels out in the total ionic equations so you know in a case chemical reactions are described with chemical equations which symbolically present the starting materials end products

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water H^+ aq OH^- aq H_2O l the spectator ions in KOH aq and HNO_3 aq are K^+ aq and NO_3^- aq how to identify and list the

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reactions and products $\text{Pb}^{2+}(\text{aq}) + 2\text{Cl}^{-}(\text{aq}) \rightarrow \text{PbCl}_2(\text{s})$ complete ionic equation an equation written with all soluble strong electrolytes shown as ions

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web the $\text{Ag}^{+}(\text{aq})$ and $\text{Cl}^{-}(\text{aq})$ ions become $\text{AgCl}(\text{s})$ but the $\text{Na}^{+}(\text{aq})$ ions and the $\text{NO}_3^{-}(\text{aq})$ ions stay as $\text{Na}^{+}(\text{aq})$ ions and $\text{NO}_3^{-}(\text{aq})$ ions these two ions are examples of spectator ions ions that do nothing in the overall course of a chemical reaction

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web jun 2 2023 the spectator ions are cut out from a reaction hence the net ionic equation for the above reaction would be $2\text{Cl}^{-}(\text{aq}) + \text{Cu}^{2+}(\text{aq}) \rightarrow \text{CuCl}_2(\text{s})$ these spectator ions are not actively involved in the reaction the existence of spectator ions affects the Debye length