

**INDORE PUBLIC SCHOOL**  
**MID TERM EXAM PRACTICE PAPER**

Class-XI

**SUBJECT – CHEMISTRY (THEORY)**

**Time : 3 Hrs**

**M.M. : 70**

**General instructions: -**

- (i) This question paper contains 26 questions. Write answers neatly and legibly. All questions are compulsory. This question paper has five sections: Section A, Section B, Section C, Section D and Section E.
- (ii) Question no. 1 to 5 are very short answer questions and carry 1 mark each.
- (iii) Question no. 6 to 10 are short answer questions and carry 2 marks each.
- (iv) Question no. 11 to 22 are also short answer questions and carry 3 marks each.
- (v) Question no. 23 is value based question and carry 4 marks.
- (vi) Question no. 24 to 26 are long answer questions and carry 5 marks each.
- (vii) Use log tables, if necessary. Use of calculators is not allowed.

**SECTION-A [5 Questions (1 Mark Each), Total Marks: 5]**

- Q.1** How many numbers of atoms are present in 1.4 gram of nitrogen gas?
- Q.2** Write the electronic configuration of Chromium and Copper? (At. No. are 24, 29 respectively)
- Q.3** What is the  $n_{\text{factor}}$  of  $\text{KMnO}_4$  in acidic medium?
- Q.4** Define ionization energy?
- Q.5** Give hybridisation of  $\text{XeO}_2\text{F}_2$  and  $\text{CO}_3^{2-}$ .

**SECTION-B [5 Questions (2 Marks Each), Total Marks: 10]**

- Q.6** Nitrogen and Hydrogen gets combined to give Ammonia 0.5 mole of  $\text{N}_2$  and 2.0 moles of  $\text{H}_2$  are taken initially. Find:
  - i) What is limiting reagent?
  - ii) Weight of  $\text{NH}_3$  gets formed?
- Q.7** Define Aufbau rule and Pauli's Exclusion Principle.
- Q.8** What are the reasons behind diagonal relationship and write the elements which shows the diagonal relationship?
- Q.9** Among F and Cl, which one have the higher electron affinity. Give reason for your answer.
- Q.10** Which is more basic,  $\text{NH}_3$  or  $\text{PH}_3$ ? Give reason for your answer.

**SECTION-C [12 Questions (3 Marks Each), Total Marks: 36]**

- Q.11**
  - a) Define Heisenberg Uncertainty Principle?
  - b) Minimum number of electron required to form a bond are?
- Q.12** Give the hybridisation and structure of  $\text{PCl}_3\text{F}_2$ .
- Q.13** The magnetic moment of an element is 3.8 B.M. ( $Z = 27$ ).
  - a) No. of unpaired electrons are.
  - b) Write electronic configuration of the element.
- Q.14** A compound contains C = 71.23%, H = 12.95% and O = 15.81%. What is the empirical formula of the compound?
- Q.15** Define Molarity (M), Molality (m) and Normality (N).

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