

Name: _____
Period: _____

Fall Semester Review – IPC

IPC Chemistry Review

Fill in the table at the right from your periodic table.

Element	Chemical Symbol	Atomic #	Valence Electrons	Oxidation #	# of Protons	Atomic Mass	Metal or Nonmetal	Group or Luster Electrons	# of Full Electron Levels
Oxygen									
Argon									
Calcium									
Nitrogen									
Potassium									
Sulfur									
Helium									

Find the molar mass of the following compounds

Calcium oxide (CaO)

Water

Classify (ionic, covalent or polyatomic) and name the following compounds.

- CO₂ _____
- N₂O _____
- NaClO₂ _____

Draw the Lewis Dot Diagrams for

- Carbon
- Sodium

Make balanced ionic compounds from:

- Li + O _____
- Be + NO₃⁻ _____
- Na + Ar _____
- Ca + O _____

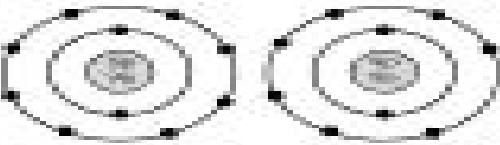
With dot diagrams draw the covalent bond of O₂

With dot diagrams the covalent bond of CO₂

1. Valence electrons are the outermost electrons of an atom that are involved in chemical bonding? True/False

2. Lithium has _____ valence electrons, will (lose or gain) _____ electrons and become (positive or negative). This is why lithium's oxidation number is _____.

3. A metal and non-metal will form a _____ compound. Two non-metals form a _____ compound.



Which one is the ion?

Give the ion notation for the ion.

Atom A
Atom B

Matching:

- Ion
- Element
- Isotope
- Oil

If you change the number of protons, you change the _____.

If you change the number of neutrons, you change the _____.

If you change the number of electrons, you change the _____.

Every 3,000 miles you should change a car's _____.

Matching:

- Unit of Conversion of Mass
- Meter
- Centimeter
- Liter
- Gram

Unit of length: equal to 3.2 ft.

Mass is neither destroyed or created in chemical reactions.

Unit of mass: about 1 dollar bill.

1/100th of mass: width of pinky finger.

Unit of volume: just bigger than a quart.

Matching:

- Solid
- Liquid
- Gas
- Oil

Molecules that are tightly packed and retain their shape and size.

Molecules that bounce off of each other, can be compressed, and take the shape of their container.

Molecules that can move (slide) over each other and have a definite size (volume), but not shape and can not be compressed.

Matching:

- Mixture
- Compound
- Element
- Matter
- Heterogeneous
- Homogeneous

A substance made up of two or more atoms that must be separated by chemical means.

Something made up of 2 or more different atoms that can be separated by physical means.

A substance that is pure and made up of only one type of atom.

Anything that has mass and takes up space.

A mixture that is the same throughout.

A mixture that is different throughout.

Ipc Chemistry Fall Semester Review Answers

JE Gale



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