- A <u>chemical bond</u> forms when 2 or more atoms rearrange <u>valence</u> <u>electrons</u> to increase <u>stability</u>.
- ionic bond forms when valence <u>clectrons</u> are <u>transferred</u>
 from one atom to another
 - cation atom <u>loses</u> electrons to become <u>+</u> charged
 - anion atom <u>sains</u> electrons to become ___ charged
 - In ionic compounds the ions are arranged in a <u>crystal</u> <u>lattice</u>

 Strong forces hold the ions together.

Properties of ionic compounds:

- high melting and boiling points
- hard not easily crushed
- conduct electricity when <u>melted</u> or

 dissolved because the ions are free to <u>move</u>
- covalent bond electrons are shared forming molecules
 - Covalent compounds have <u>weaker</u> forces holding the <u>molecules</u> together.

Properties of covalent compounds:

- lower melting and boiling points
- Many covalent compounds are <u>volatile</u> liquids or gases.
- soft ___ easier to __crush ______
- are not <u>conductors</u> of electricity
- electronegativity property that tells how strong an atom's <u>attraction</u> is for <u>electrons</u>
- Since oxygen has a <u>higher</u> electronegativity than hydrogen, oxygen holds onto shared electrons <u>tighter</u>, giving the oxygen a <u>partial</u> negative charge and the hydrogen a partial <u>positive</u> charge.

CHEMISTRY: A Study of Matter

Ionic Bonds And Notetaking Guide Answer

Christian Drosten

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