

Name Answer Key Hour _____

Chapter 4 Review (Honors Chemistry)

Use the given number of neutrons and electrons to fill in the chart and determine whether the atom is...

Common: all information is exactly like the information on the periodic table.

Ion: the electron number differs from the proton number, but all other information is the same.

Isotope: the mass of the atom is different than the most common version.

	ELEMENT	ATOMIC NUMBER	MASS	PROTON NUMBER	NEUTRON NUMBER	ELECTRON NUMBER	ISOTOPE, ION OR COMMON
1.	Sc	21	45	21	24	18	Cation, Sc^{3+}
2.	K	19	42	19	23	19	Isotope
3.	Cobalt	27	59	27	32	29	Anion, Co^{2-}
4.	S	16	32	16	16	18	Anion, S^{2-}
5.	Argon	18	40	18	22	18	Common
6.	Uranium	92	238	92	146	92	Common
7.	Ca	20	43	20	23	20	Isotope
8.	P	15	31	15	16	18	Anion, P^{3-}
9.	Pb	82	225	82	143	82	Isotope
10.	Zinc	30	65	30	35	30	Common
11.	Aluminum	13	30	13	17	13	Isotope
12.	Potassium	19	39	19	20	18	Cation, K^{+}
13.	Ce	58	142	58	84	58	Isotope
14.	Sodium	11	23	11	12	10	Cation, Na^{+}
15.	Chlorine	17	35	17	18	18	Anion, Cl^{-}
16.	Radon	86	222	86	136	86	Common
17.	Mg	12	22	12	10	12	Isotope
18.	Boron	5	11	5	6	5	Common
19.	Nickel	28	59	28	31	26	Cation, Ni^{2+}
20.	Bismuth	83	209	83	126	83	Common
21.	Iridium	77	196	77	119	77	Isotope
22.	Hg	80	201	80	121	78	Cation, Hg^{2+}
23.	Titanium	22	50	22	28	22	ISOTOPE

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Victor M. Corman



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