

Equilibria Involving A Weak Base

You have 0.010 M NH_3 . Calc. the pH.



$$K_b = 1.8 \times 10^{-5}$$

Step 1. Define equilibrium concs. in ICE table

	$[\text{NH}_3]$	$[\text{NH}_4^+]$	$[\text{OH}^-]$
initial	0.010	0	0
change	-x	+x	+x
equilib	0.010 - x	x	x

Honors Chemistry Equilibrium

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1. Understanding the eBook Honors Chemistry Equilibrium
 - The Rise of Digital Reading Honors Chemistry Equilibrium
 - Advantages of eBooks Over Traditional Books
2. Identifying Honors Chemistry Equilibrium
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Honors Chemistry Equilibrium
 - User-Friendly Interface
4. Exploring eBook Recommendations from Honors Chemistry Equilibrium
 - Personalized Recommendations
 - Honors Chemistry Equilibrium User Reviews and Ratings
 - Honors Chemistry Equilibrium and Bestseller Lists

5. Accessing Honors Chemistry Equilibrium Free and Paid eBooks
 - Honors Chemistry Equilibrium Public Domain eBooks
 - Honors Chemistry Equilibrium eBook Subscription Services
 - Honors Chemistry Equilibrium Budget-Friendly Options
6. Navigating Honors Chemistry Equilibrium eBook Formats
 - ePub, PDF, MOBI, and More
 - Honors Chemistry Equilibrium Compatibility with Devices
 - Honors Chemistry Equilibrium Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Honors Chemistry Equilibrium
 - Highlighting and Note-Taking Honors Chemistry Equilibrium
 - Interactive Elements Honors Chemistry Equilibrium
8. Staying Engaged with Honors Chemistry Equilibrium
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Honors Chemistry Equilibrium
9. Balancing eBooks and Physical Books Honors Chemistry Equilibrium
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Honors Chemistry Equilibrium
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Honors Chemistry Equilibrium
 - Setting Reading Goals Honors Chemistry Equilibrium
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Honors Chemistry Equilibrium
 - Fact-Checking eBook Content of Honors Chemistry Equilibrium
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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