

Half-Modern Chemistry Review

CHAPTER 11: CHEMICAL BONDING

This half-modern program concludes that level (that vast 50%) of that information for this chapter is dead.

Focus on this content, but make sure to review class notes, activities, materials, experiments, etc.

If your study this document and DISCOVERED sites, you should at least be able to PLUG that level.

Now that there will be recall, examples, and/or application of this content, now

OUTCOMES

- Communicate with precision the understanding/properties learned (p2 c2)
- Communicate understanding learned to teachers and/or peers (p2 c2)
- EL-2 Give the Lewis dot notation, illustrate the placement bond on the drawing of molecules (p2 d, p2c)
- EL-2 List important physical properties of some compounds; name and illustrate molecules in ions (p2 d, p2c)

EL-1: INTRODUCTION TO CHEMICAL BONDING

• Introductory

- a) chemical bond → the attractive force that holds atoms or ions together
- a) ionic bonding → a force that attracts particles that have been able to attract, which demonstrates a mutual attraction, and can be
- a) covalent bonding → a bond formed when atoms share electrons with one or more pairs of electrons.
- a) intermolecular bonding → a bond formed by attractive forces between molecules and especially between the bonds between atoms.
- a) gas → substance, a molecule or particle that particles and molecules change and change and
- a) gas-liquid bond → a substance bond in which a gas and molecules change by being at some or both ends of the spectrum.

• Chapter Highlights

- a) ionic atoms are chemically changed by other atoms. The most major types of chemical bonding are ionic, covalent, and metallic.
- a) In general, atoms of metals, bond covalently with atoms of nonmetals, atoms, or metals bond covalently with each other, and atoms of nonmetals bond covalently with a metal atom.

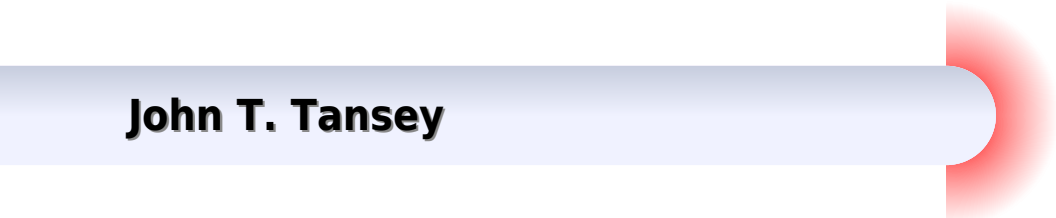
EL-2: COVALENT BONDING AND MOLECULAR COMPOUNDS

• Introductory

- a) molecule → a group of atoms that are held together by chemical bonds; a molecule is the smallest unit of matter that can exist by itself and retain all of its chemical properties
- a) molecular compound → a chemical compound whose simplest unit, and molecules, is chemical formula
- a) chemical formula → a combination of chemical symbols and numbers that represents a substance
- a) molecular formula → a chemical formula that shows the number and kinds of atoms in a molecule, but not the arrangement of the atoms.

Holt Chemistry Chapter 6 Review Answers

John T. Tansey



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