

4NB – VIDHATA

Date: 30/09/2014

DURATION : 2 Hrs 30 Min.

MAX MARKS : 75

INSTRUCTIONS TO THE STUDENTS :

- All questions are compulsory.
- Figures to the right indicate the marks.
- Draw neat sketches / diagrams / graphs / Flow sheets whenever necessary.
- Use of simple / scientific calculator / log tables permitted.

- A.] Answer Any Two of the following.

 - Predict the shape of following molecules :
 BeCl_2 , NH_3 , PCl_5 . 06
 - State and explain Ionization Energy. 03
 - What is Pauli's Rule? 03
 - Write short note on : Covalent Bond. 03
 - What are limitations of VSEPR theory? 03
 - What is Electronegativity? How does it vary in period and group in the periodic table? 03
 - Explain the Heisenberg's principle of Uncertainty? 03

B.] Answer Any Two of the following.

 - What is lone pair effect? How does it affect the bond angles in molecules? 04
 - What are applications of VSEPR theory? 04
 - Explain: BF_3 is planar whereas NF_3 is not planar. 04
 - Discuss the Rutherford's atomic model. 04
- A.] Answer Any Two of the following.

 - Define the following :-
Hydrogen Bond, Ionic Bond, Dipole Moment. 03
 - Calculate the formal charge on the atom in the following :
Nitrogen in Ammonium ion
Oxygen in Water
Nitrogen in Nitromethane 03
 - What are Dry Test and Wet Test? Give any two examples for each test. 06
 - What are types of Arrows and their significance? 03
 - Why Formic Acid is stronger than Acetic Acid? 03
 - Explain the structure of following functional groups:
Ether, Cyanide, Ketone. 06

B.] Answer Any Two of the following.

 - What is Inductive Effect? Explain its types. 04
 - What is Semi-micro Analysis? 04
 - What is Formal Charge? Explain any two examples. 04
 - Explain sp^3 hybridization for Nitrogen and Oxygen. 04

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